Antacid Analysis and Titration – Lab Report Assistant Exercise 1: Back Titration of Antacid Neutralization

Data Table 1. Antacid Neutralization Data

| Mass of Crushed Antacid (g) | |
|--------------------------------|------|
| Concentration of HCI (M) | 1M |
| Volume HCI (mL) | 5 mL |
| Concentration of NaOH (M) | 1M |
| Initial NaOH Volume (mL) | |
| Final NaOH Volume (mL) | |
| Total Volume of NaOH Used (mL) | |

Data Table 2. Experimental Results

| HCI available for neutralization (g): | 0.3646 (g) | |
|---|----------------------------|--|
| Moles of NaOH required to reach stoichiometric point: | 0.0039 (mols) 0.222 (g) | |
| HCl neutralized by antacid: | | |
| HCl neutralized per gram of antacid: | 0.445 (g) | |

Questions:

- A) If an antacid tablet weighed 1.6 grams, how many moles of gastric acid (HCI) would it neutralize? Use the results obtained in Data Tables 1 and 2 to explain and quantify your answer.
- B) If you performed this experiment with a titrant of 0.5M NaOH, would you expect your results in Data Table 2 to change or stay the same? Explain your answer.
- C) The reaction that occurred when the antacid mixed with the HCl resulted in an additional product (besides a salt and water). Did you see evidence of this product? Describe the experimental evidence you witnessed that supports the formation of the additional product.

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