

## **Internal/Industrial Attachment Report format**

### **1. Cover/ Title page**

*Sample Title page*



## **ENGINEERING INDUSTRIAL/INTERNAL ATTACHMENT REPORT**

***INTERNAL/INDUSTRIAL ATTACHMENT CONDUCTED AT MURANG'A UNIVERSITY  
LABS & WORKSHOPS / KENYA POWER-THIKA STATION***

**SUBMITTED**

**TO**

**MURANG'A UNIVERSITY OF TECHNOLOGY**

**DEPARMENT OF ELECTRICAL & ELECTRONIC ENGINEERING**

**BY**

***RUTH NJERI- 2<sup>nd</sup> YEAR***

***BTECH ELECTRICAL & ELCTRONIC***

***ENGINEERING REG. NO: EN200/0205/2013***

***APRIL 2020, ACADEMIC YEAR 2019/2020***

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hot object is 60 c the specific heat of copper and water are 0 09 cal gm c and 1 cal gm c respectively

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temperature change plugging in the values given in the problem you get  $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 12500 \text{ J}$   
 $550 \text{ J} = \Delta h q_{\text{water}} = 550 \text{ J}$

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**heat capacity and calorimetry practice khan academy** - Jun 04 2023

web choose 1 answer the final temperature is closer to t<sub>1</sub> than to t<sub>2</sub> a the final temperature is closer to t<sub>1</sub> than to t<sub>2</sub> the final temperature is exactly halfway between t<sub>1</sub> and t<sub>2</sub> b the final temperature is exactly halfway between t<sub>1</sub> and t<sub>2</sub> the final temperature is closer to t<sub>2</sub> than to t<sub>1</sub> c

*calorimetry practice problems chemistry steps* - Sep 07 2023

web in this set of practice questions we will go over the main types of questions on calorimetry including the heat capacity the heat of reaction finding the final temperature of a mixture constant pressure calorimetry and constant volume calorimetry a few important concepts and formulas you will need to solve these questions

*calorimetry worksheet laney college* - Nov 28 2022



web calorimetry worksheet 1 if 0.315 moles of hexane  $C_6H_{14}$  is combusted in a bomb calorimeter containing 5.65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 5.54 °C the specific heat capacity of water is 4.184 J/g °C h m s t h 5.650 grams  $H_2O$  4.184 J/g °C 5.54 °C h 1310 kJ

*calorimetry problems 1* - Feb 17 2022

web 6 a 218 g sample of steam at 121°C is cooled to ice at 14°C find the change in heat content of the system 7 in going from ice at 34°C to steam at 138°C a sample of water absorbs  $1.41 \times 10^5$  J find the mass of the sample answers 1 2 68 x 104 J 2 6 23 x 104 J 3 1 71 x 104 J 4 1 11 x 105 J 5 2 11 x 105 J 6

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