

DEVOIR DE CONTROLE N°2

Date: 10 -02-2020

Durée : 2H

- le sujet comporte 2 exercices de chimie et 2 exercices de physique.
- On donnera l'expression littérale avant de passer à l'application numérique.
- L'utilisation de la calculatrice scientifique est autorisée.



CHIMIE :(9 POINTS)

Exercice n°1 :

Un acide carboxylique (A) de masse molaire $M = 74 \text{ g.mol}^{-1}$.

- Déterminer la formule semi développée et le nom de (A).
- On dissout une masse $m=0,74\text{g}$ de (A) dans l'eau distillée afin de préparer 100mL d'une solution (S) d'acide de concentration molaire C. Le pH de cette solution est égal à 3.
 - Calculer la concentration molaire C.
 - Comparer C et $[H_3O^+]$. Ecrire son équation de dissociation dans l'eau.
 - Déduire la molarité des espèces chimiques présentes dans la solution (sauf l'eau).
On rappelle qu'à 25°C : $[H_3O^+][OH^-]=10^{-14}$
- On fait réagir un volume $V^*=50\text{mL}$ de la solution (S) avec $m^*=0,1308\text{g}$ de zinc, il se dégage un gaz qui provoque une légère détonation en présence d'une flamme.
Quel est le gaz dégagé ? Calculer son volume.
- On fait réagir 5 moles de (A) avec 5 moles d'alcool (B) qui donne à la fin de la réaction 3 moles d'un produit (C) formé par 6 atomes de carbone de plus 3 moles d'eau.
 - Déterminer la formule brute de l'alcool (B) ainsi que la formule semi-développé de (C) et son nom.
 - Dégager de ce qui précède la propriété de cette réaction.

0,75B

0,75B

0,75B

0,75B

1C

1C

0,5B

Exercice n°2 :

La combustion d'une masse $m_1=2\text{g}$ d'un alcool (A) $C_nH_{2n+2}O$ fournit $m_2=5\text{g}$ de dioxyde de carbone.

- Ecrire l'équation générale de la réaction de combustion.
- Montrer que la masse molaire de l'alcool (A) peut s'écrire $M=17,6n$.
- Déduire alors la formule brute de (A)
- L'oxydation ménagée de (A) par le bichromate de potassium ($2K^+$, $Cr_2O_7^{2-}$) en milieu acide donne un composé (B) qui réagit avec le 2,4 DNPH mais ne réagit pas avec le réactif de Schiff.
 - Préciser la fonction chimique de (B) ?
 - Déduire alors tous les isomères possibles de (A).
- Ecrire l'équation de la réaction d'oxydation ménagée en utilisant les formules brutes.
- La déshydratation intramoléculaire de (A) donne un seul composé organique (C). Identifier les composés organiques (A), (B) et (C).

0,5A

0,5 C

0,5B

0,25A

0,75A

1C

1C

On donne g. mol^{-1} : $M(C)=12$; $M(O)=16$; $M(H)=1$; $M(Zn)=65,4$ et $V_m=24 \text{ L.mol}^{-1}$



Devoir De Controle N 2 3 M Exercice1

M Lipman



Devoir De Controle N 2 3 M Exercice1 :

Math Max - Première enseignement de spécialité - Cours complet, exercices et devoirs corrigés - Nouveaux programmes
Sébastien Krief-Détraz, 2021-01-05 Cette seconde édition conforme aux nouveaux programmes contient un cours complet avec des exemples des remarques et des conseils des centaines d'exercices et devoirs tous corrigés en détail de difficulté croissante couvrant tout le programme et même plus des cahiers transversaux de logique et d'algorithmique des activités d'approche de révision et de synthèses extras pour anticiper sur les années venir des exercices en anglais pour enrichir la langue des corrections rédigées comme l'exigent les enseignants une approche testée et validée auprès des élèves Tout ce qui est nécessaire pour asseoir ses bases réussir son année et bien préparer la suivante Mais Math Max c'est aussi des données historiques des conseils de lectures des touches culturelles des exercices d'actualités des noncés fantasques des contextes farfelus des jeux de mots presque amusants

Spé Physique-Chimie. Première Sophie Auvray, 2025-07-15 Vous êtes en Première et vous cherchez un livre de Physique Chimie pour vous entraîner aux devoirs sur table et avoir la note maximum Voici l'ouvrage qu'il vous faut Vous apprendrez à crypter un nonc le relier aux notions apprises en cours et vous organiser lors de la correction pour gagner en rapidité et efficacité La première partie du livre regroupe tous les devoirs de l'année sachant qu'un devoir correspond aux connaissances d'une semaine de cours vues en classe La deuxième partie rassemble tous les corrigés détaillés et expliqués Les plus Une table thématique pour travailler les exercices dans le désordre De nombreux conseils de l'enseignant Des explications sur le déroulement de l'année Une méthode de travail Les astuces pour réussir son contrôle Les mots et les formules connaître par cœur En bonus Chaque exercice est chronométré avec une note sur 20 pour être au plus proche des conditions réelles du devoir sur table

Recueil général des lois et des arrêts, 1890 *BLOC 2 - Gestion de projet - BTS SAM 1re et 2e années - Éd. 2018 - Manuel FXL* Nadège Cadet-Snelder, Nadia Chougrani-Nief, Brigitte Denez, Sylvie Dufour, Chantal François-Gustin, Marie-Pierre Guilbert-Varlet, 2018-04-25 Bloc 2 Gestion de projet 1 carnet de bord en double ouvrage et propre chaque matière rappelle aux étudiants les outils et prérequis essentiels Une approche par situations professionnelles variées permet l'étudiant de se projeter en tant que support d'un ou de plusieurs managers dans différents types d'organisations Dans chaque chapitre un objectif atteindre est détaillé travers 2 à 4 missions que les étudiants pourront réaliser un travail concret qui s'appuie sur des annexes contextuelles et notionnelles 1 page de synthèse donne l'essentiel retenir 1 partie entraînement propose des exercices progressifs En fin d'ouvrage des fiches pratiques avec les outils nécessaires la réalisation des missions Des tutoriels vidéo les complètent pour une plus grande autonomie des étudiants Ce manuel est enrichi de ressources numériques FOUCHER CONNECT Des pictogrammes indiquent au fil des pages des liens vers des vidéos des tutoriels des schémas de synthèse animés en accès direct sur www.foucherconnect.fr Profitez de l'offre MON COACH BESCHERELLE en partenariat avec la collection Atouts Compétences Une application pour progresser en orthographe et en rédaction grâce à un entraînement 100% personnalisé Cliquez ICI pour accéder à l'offre

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