

Lab 12C

Determination of a Solubility Product Constant

Purpose: To prepare and investigate solutions of Pb^{2+} and I^- , determining solubility concentrations, and then determine corresponding values. Calculate also obtain an approximate value of K_{sp} for PbI_2 at different temperatures.

Safety: Lead(II)nitrate is very toxic. Handle immediately if spilled!

Procedure:

- a) When is a solution said to be saturated?

A solution is saturated when the rate of dissolving is equal to the rate of crystallization, meaning that no more solute will be dissolved.

- b) What quantities are used to represent the two concentrations for use in a K_{sp} expression?

The coefficient of each product in balanced equation divided by that number (e.g. Ca^{2+} coefficient).

- c) What happens when the solubility [] of ions substituted into a K_{sp} expression equals the actual K_{sp} ?

A mixture that solution is then kept saturated so that the fully saturated and some will fall out of the solution and crystallizing solute.

Table 1:

Test Tube	1	2	3	4	5	6
Vol. H_2O (mL)	10.0	10.0	10.0	10.0	10.0	10.0
Prep. Conc. (mol/L)						
Vol. $\text{Pb}(\text{NO}_3)_2$ solution (mL)	10.0	10.0	10.0	10.0	10.0	10.0
Vol. KI solution added (mL)	0.0	1.0	2.0	3.0	4.0	5.0
Vol. $\text{Pb}(\text{NO}_3)_2$ solution (mL)	10.0	10.0	10.0	10.0	10.0	10.0
Vol. KI solution added (mL)	0.0	1.0	2.0	3.0	4.0	5.0
Prep. Conc. of Pb^{2+} (mol/L)	0.01	0.01	0.01	0.01	0.01	0.01
Prep. Conc. of I^- (mol/L)	0.0	0.001	0.002	0.003	0.004	0.005
Temperature at which precipitation observed ($^{\circ}\text{C}$)						

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Answers

Jianjun Gao



Determination Of A Solubility Product Constant Lab 12c Answers:

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