

Ideal Gas Law

Chem Worksheet 14-4

Name _____

The ideal gas law is an equation that relates the volume, temperature, pressure and amount of gas particles to a constant. The ideal gas constant is abbreviated with the variable R and has the value of $0.0821 \text{ atm}\cdot\text{L}/\text{mol}\cdot\text{K}$. The ideal gas law can be used when three of the four gas variables are known. When using this equation it is important that the units for **pressure** are atmospheres (atm), **volume** is in liters (L), and **temperature** is converted to kelvins (K). The **amount** of gas is measured in units called moles (mol).

USEFUL EQUATIONS

$$PV = nRT$$

$$R = 0.0821 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}}$$

$$T_K = T_C + 273$$

$$1 \text{ cm}^3 = 1 \text{ mL}$$

$$1 \text{ L} = 1000 \text{ mL}$$

$$1.00 \text{ atm} = 101300 \text{ Pa}$$

$$1.00 \text{ atm} = 101.3 \text{ kPa}$$

$$1.00 \text{ atm} = 760 \text{ mmHg}$$

$$1.00 \text{ atm} = 760 \text{ torr}$$

$$1.00 \text{ atm} = 14.7 \text{ psi}$$

Unknown	Equation	Known Variables
pressure	$P = \frac{nRT}{V}$	amount, temp., volume
volume	$V = \frac{nRT}{P}$	amount, temp., pressure
temperature	$T = \frac{PV}{nR}$	pressure, volume, amount
amount	$n = \frac{PV}{RT}$	pressure, volume, temp.

example

The pressure exerted by 2.8 moles of argon gas at a temperature of 85°C is 420 torr. What is the volume of this sample?

- list the variables: $P = 420 \text{ torr}$ $V = ?$ $n = 2.8 \text{ mol}$ $R = 0.0821 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}}$ $T = 85^\circ\text{C}$
- convert the variables: $\frac{420 \text{ torr}}{1} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 0.553 \text{ atm}$ $T = 85^\circ\text{C} + 273 = 358 \text{ K}$
- substitute into the equation: $V = \frac{nRT}{P}$ $V = \frac{(2.8 \text{ mol})(0.08206 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}})(358 \text{ K})}{0.553 \text{ atm}} = 82 \text{ L}$

Solve the following problems.

1. A tank contains 115 moles of neon gas. It has a pressure of 57 atm at a temperature of 45°C . Calculate the volume of the tank.
2. A scuba tank has a pressure of 195 atm at a temperature of 10°C . The volume of the tank is 350 L. How many moles of air are in the tank?
3. A helium-filled balloon has a volume of 208 L and it contains 9.95 moles of gas. If the pressure of the balloon is 1.26 atm, determine the temperature in Celsius degrees.
4. A tank of oxygen has a volume of 1650 L. The temperature of the gas inside is 35°C . If there are 9750 moles of oxygen in the tank what is the pressure in PSI?
5. A canister of acetylene has a volume of 42 L. The temperature of the acetylene is 305 K and the pressure is 780 torr. Determine the amount (moles) of gas in the canister.
6. Calculate the volume of a CO_2 cartridge that has a pressure of 850 PSI at a temperature of 21°C . The cartridge contains 0.273 mol of CO_2 .
7. A tank contains 2500 L of argon gas. The pressure is 13790 kPa and the temperature is 25°C . How many moles of argon are in the tank?

Ideal Gas Law Chem 14 4 Answers

**Barron's Educational Series, Neil D.
Jespersen, Pamela Kerrigan**

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Chemical Thermodynamics at a Glance H. Donald Brooke Jenkins, 2008-04-30 Chemical thermodynamics considers the energy transformations which drive or which occur as a result of chemical reactions It is a central discipline of chemistry and chemical engineering allowing prediction of the direction of spontaneous chemical change and the position of chemical equilibrium in any reacting system Being grounded in maths it is often perceived as a difficult subject and many students are never fully comfortable with it Chemical Thermodynamics at a Glance provides a concise overview of the main principles of Chemical Thermodynamics for students studying chemistry and related courses at undergraduate level Based on the highly successful and student friendly at a Glance approach the information is presented in integrated self contained double page spreads of text and illustrative material The material developed in this book has been chosen to ensure the student grasps the essence of thermodynamics so those wanting an accessible overview will find this book an ideal source of the information they require In addition the structured presentation will provide an invaluable aid to revision for students preparing for examinations

Physical Chemistry Horia Metiu, 2006-02-21 This is a new undergraduate textbook on physical chemistry by Horia Metiu published as four separate paperback volumes These four volumes on physical chemistry combine a clear and thorough presentation of the theoretical and mathematical aspects of the subject with examples and applications drawn from current industrial and academic research By using the computer to solve problems that include actual experimental data the author is able to cover the subject matter at a practical level The books closely integrate the theoretical chemistry being taught with industrial and laboratory practice This approach enables the student to compare theoretical projections with experimental results thereby providing a realistic grounding for future practicing chemists and engineers Each volume of Physical Chemistry includes Mathematica and Mathcad Workbooks on downloadable resources Metiu's four separate volumes Thermodynamics Statistical Mechanics Kinetics and Quantum Mechanics offer built in flexibility by allowing the subject to be covered in any order These textbooks can be used to teach physical chemistry without a computer but the experience is enriched substantially for those students who do learn how to read and write Mathematica or Mathcad programs A TI 89 scientific calculator can be used to solve most of the exercises and problems Mathematica is a registered trademark of Wolfram Research Inc Mathcad is a registered trademark of Mathsoft Engineering Education Inc

Transport Phenomena Problem Solver, 1984

Fundamentals of Gas Dynamics Howard W. Emmons, 2015-12-08 Volume II of the High Speed Aerodynamics and Jet Propulsion series The series which stress the more fundamental aspects of the various phenomena that make up the broad field of aeronautical science The aerodynamicist and gas dynamicist will find both the classical and the important new concepts of gas dynamics presented in an informative and stimulating manner Specialists in the study of gas dynamics have contributed Sections as follows H S Tsien The Equations of Gas Dynamics L Crocco One Dimensional Treatment of Steady Gas Dynamics A Kantrowitz One Dimensional Treatment of Nonsteady Gas Dynamics W

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Physical and Chemical Equilibrium for Chemical Engineers Noel de Nevers,2012-03-20 This book concentrates on the topic of physical and chemical equilibrium Using the simplest mathematics along with numerous numerical examples it accurately and rigorously covers physical and chemical equilibrium in depth and detail It continues to cover the topics found in the first edition however numerous updates have been made including Changes in naming and notation the first edition used the traditional names for the Gibbs Free Energy and for Partial Molal Properties this edition uses the more popular Gibbs Energy and Partial Molar Properties changes in symbols the first edition used the Lewis Randal fugacity rule and the popular symbol for the same quantity this edition only uses the popular notation and new problems have been added to the text Finally the second edition includes an appendix about the Bridgman table and its use **2024-25 NCERT Class 11th**

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