

# Calculations in IB Chemistry SL and HL

Below is a list of the calculations are expected to be able to do in IB chemistry. This needs committing to memory, and each type of calculation will need practice to master. Good luck!

Calculations shaded in grey are only required at HL

## Quantitative Chemistry

Purpose	The Calculation	Notes
1. Calculating a quantity in moles from a number of particles	$n = \frac{N}{L}$ <p>n = the quantity in moles N = the number of particles L = Avogadro's constant, <math>6.02 \times 10^{23}</math></p>	<ul style="list-style-type: none"> <li>When calculating numbers of atoms within molecules, multiply the number of particles (N) by the number of atoms in the formula</li> <li>To find out numbers of particles, rearrange to: <math>N = n.L</math></li> </ul>
2. Determine relative molecular or formula mass, $M_r$	$M_r = \sum (\text{number of atoms} \cdot \text{atomic mass})$	<ul style="list-style-type: none"> <li><math>M_r</math> has no unit as it is a relative value</li> <li>To calculate molar mass, <math>M_{mol}</math> just stick a 'g' for grams on the end</li> </ul>
3. Calculate a quantity in moles, from a mass of a substance	$n = \frac{m}{M_{mol}}$ <p>n = the quantity in moles m = the mass of substance you are given in grams <math>M_{mol}</math> = the molar mass of the substance</p>	<ul style="list-style-type: none"> <li>To determine the mass of a given number of moles of a substance use: <math>m = M_{mol} \cdot n</math></li> <li>To determine the molar mass of a given mass of a substance: <math>M_{mol} = \frac{m}{n}</math></li> </ul>
4. Determine empirical formula from % composition by mass	<ol style="list-style-type: none"> <li>Divide each % by the atomic mass</li> <li>Divide each answer to Step 1 by the smallest answer to Step 1</li> <li>Multiply all answers to Step 1 to remove any obvious fractions                             <ol style="list-style-type: none"> <li>If there is a '.5' multiply everything by 2</li> <li>If there is a '.3' multiply everything by 3 etc</li> </ol> </li> </ol>	
5. Determine molecular formula from empirical formula	$F_{mol} = \frac{M_r}{m(F_e)} \cdot F_e$ <p><math>F_{mol}</math> = molecular formula <math>M_r</math> = relative molecular mass <math>F_e</math> = empirical formula <math>m(F_e)</math> = empirical formula mass</p>	<ul style="list-style-type: none"> <li>This and the previous calculation are often combined together in exam questions</li> </ul>
6. Use mole ratios to determine the number of moles of B that can be made from A	$n(B) = n(A) \cdot \frac{\text{number of B in equation}}{\text{number of A in equation}}$	<ul style="list-style-type: none"> <li>The second term in this equation is the mole ratio</li> <li>You must use a fully balanced equation</li> <li>This is the central step in many stoichiometry calculations</li> </ul>
7. Calculate theoretical yield	<ol style="list-style-type: none"> <li>Use <b>Calculation 6</b> to determine the expected quantity of product in moles</li> <li>Use a rearranged <b>Calculation 3</b> to determine the expected mass.</li> </ol>	

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**Daniela Niemeyer**



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