

# **Experiment 3: Basic Protonation, Propanolamine and Nucleophilicity**

Basic propanolamine (see Page 1, Worksheet 234.3B - 19, pp 477-478)



The (propanolamine) equilibrium constant for the reaction above is

$$K_{\text{eq}} = \frac{\text{[alkyl ammonium] [H}_2\text{O]}}{\text{[basic amine] [alcohol]}} = 10.00$$

Use the basic equilibrium procedure. The direction to go for deriving the acid distribution by collecting the concentrations of each of the species that characterize propanolamine (2-aminoethanol is propanolamine).

I. Basic Distribution (see Page 1, Worksheet 234.3B - 19, pp 477-478)

Experimental composition

	$\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$	$\rightleftharpoons$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}_2^+$	$\xrightarrow{\text{H}^+}$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{NH}_3^+$	$\rightleftharpoons$	$\text{H}_2\text{O}$
10.00000	1.00		1.00		0.1		0.1
at equilibrium	0.999		0.999		0.99		0.999
20.00000	1.00		0.999		0.1		0.1
at equilibrium	0.999		0.999		0.997		0.997
30.00000	1.00		0.999		0.1		0.1
at equilibrium	0.999		0.999		0.999		0.999

Worksheet: Worksheet 234.3B-19, Worksheet 234.3B-19, Worksheet 234.3B-19, Worksheet 234.3B-19, Worksheet 234.3B-19

II. The K<sub>eq</sub> values for the Acid-catalyzed Basic Distribution (Commonly referred to as the Reaction between Basic and pp 477-478, 478's results)



# Experiment 3 Ester Formation Preparation Of Benzocaine

**WJ Hussar**



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