

Sample Problem



Find the oxidation state of the elements in $K_2Cr_2O_7$

 	Ion	K	Cr	O	TOTAL
	Subscript	2	2	7	
	Oxidation State	+1	+6	-2	
	Sum of Ox. St.	+2	+12	-14	= 0






Prepare a table as above

- **Step 1:** List the subscripts for elements.
- **Step 2:** Apply the rules with the help of the *Periodic Table*.
 - potassium is a group one metal; its oxidation state is always +1
 - oxygen usually has an oxidation state of -2
- **Step 3:** Multiply the oxidation state of the elements by their subscripts to get the sum of the oxidation states.
- **Step 4:** Determine the sum of the oxidation states for the unknown by setting the total sum of the oxidation states to zero.
- **Step 5:** Divide the sum of the oxidation states of the unknown by its subscript to get its oxidation state.

Evan P Silberstein Oxidation Answers

Tobias Bleicker



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