In Table 4 below, [Fe³⁺] after mixing is obtained by using $M_1V_1 = M_2V_2$ This data is to be transferred to Table 5 below as the initial [Fe³⁺] (column 2).

Similarly in Table 4 of Data sheet [SCN] after mixing obtained using $M_1V_1 = M_2V_2$. This data is then to be transferred to Table 5 below as the initial [SCN] (column 3).

Table 4. Initial Concentration Data

Test Tube	[Fe ³⁺] Before Mixing	[Fe ³⁺] After Mixing ⁴ (Initial [Fe ³⁺])	[SCN ¹⁻] Before Mixing	[SCN ¹ -] After Mixing ¹ (Initial [SCN ¹ -])	
F	0.01	0.0025M	0.00113	2.835110-4	
G	0.01	0.0025M	0.00113	5.65 110-4	
н	0.01	0.00514	0.00113	2.825×10-4	
1	0.01	0.005M	0.00113	1.13 × 10 -4	

(Show calculations for initial concentration data on separate sheet)

In Table 5 below, Equilibrium [FeSCN2*] is obtained from the Calibration line using absorbance values of solutions F through I.

Equilibrium [Fe3+] = Initial [Fe3+] - Equilibrium [FeSCN2+]

Equilibrium [SCN] = Initial [SCN] - Equilibrium [FeSCN2*]

Table 5. Equilibrium Constant Calculation

Test Tube	Initial [Fe ³⁺]	Initial [SCN ¹⁻]	Equilibrium [FeSCN ²⁺] (from graph)	Equilibrium [Fe ^{3*}]	Equilibrium [SCN ¹⁻]	K.
F	0.0035	2.825x10-4				
G	0.0025	5.65×10-4		19-		
Н	0.005	2.825x10				
1	0.005	1.13210-4				Sec.

(Show calculations for Equilibrium Constant on separate sheet)

Average Equilibrium Constant

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N Noddings

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