

Lab 9 Determine of K_a of a Weak acid

INTRODUCTION

The objective of this experiment is to determine the molar mass, pK and K_a of a weak acid based on a titration curve. An acid is a compound that can donate a proton, and the strength of an acid can be determined by the compound's ability to donate a proton. The strength of a 1 acid is determined by the pH or negative log of the hydronium ion concentration. A titration curve is a graphical representation of the pH of a solution during a titration. At the equivalence point of the titration, the moles of the weak acid and the base added are equal and it corresponds with a color change of the indicator.

The equivalence point is the inflection point of the titration curve. At an inflection point, the second derivative changes from positive to negative and the first derivative has a maximum. At half of the equivalence point, the pH is equivalent to the pK_a because the molar concentration of the resulting conjugate base is equal to the molar pK_a concentration of the acid are equal.

Experimental Method:

The first step of this experiment was to conduct a test to calibrate the pH probe. This was done by placing the probe in buffer solutions of known pH. The probe was removed from the storage solution and rinsed with distilled water. The probe was placed in two known solutions, one with a pH of 10 and the other with a pH of 4 to calibrate the sensor. The probe was rinsed between each test. The second step of this experiment was to make a standardized NaOH solution.

Graphs:



Graph For Maleic Acid

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