

Quiz & Worksheet - Colligative Properties and Raoult's Law

1.3	What	happens	to vapor	pressure when	you add a	solute	to a solution?
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- It lowers the vapor pressure.
- It has no effect.
- It raises the vapor pressure.
- It causes the reaction to reach equilibrium.

2. What is osmotic pressure?

- It is the minimum pressure that must be applied to a solution to stop osmosis from happening.
- It is the maximum pressure that must be applied to a solution to stop osmosis from happening.
- c. It is the maximum pressure of the vapor over a liquid at equilibrium.
- It is the minimum pressure of the vapor over a liquid at equilibrium.

3.

9.0g of urea (a nonelectrolyte) is dissolved in 360 g water to create a solution. The temperature of the solution is 25.0°C. If the vapor pressure of water at 25.0°C is 23.80 mmHg, calculate the vapor pressure of the urea solution at 25.0°C in mmHg.

The molar mass of urea 60g/mol and water is 18g/mol.

- 0.099
- 0.2391
- 0 23.61
- 40 3.57

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