Photoelectric Effect - Problem Solving

Useful physical constants and conversions:

| $q_{electron} = -1.6\times 10^{-19}~C$ | $c=3.00\times 10^8m/s$ | $h = 6.63 \times 10^{-34} \text{ J s}$ = $4.14 \times 10^{-15} \text{ eV s}$ | $1 eV = 1.6 \times 10^{-19} J$ |
|----------------------------------------|------------------------|---------------------------------------------------------------------------------|---------------------------------|
| $m = 9.11 \times 10^{-11} kg$ | hc = 1240 eV nm | | |

Emergy relationships:

| $E_{photon} = hf = \frac{hc}{\lambda}$ | $E_{kinetic} = \frac{1}{2}mv^2$ | $\Delta E_{electric} = q \Delta V$ | $\phi = work function$ |
|----------------------------------------|--------------------------------------------------|---------------------------------------------|----------------------------|
| $KE_{max} = E_{photon} - \phi$ | qV _{stopping} = E _{photon} − φ | $hf_{\phi} = \frac{hc}{\lambda_{-}} = \phi$ | $KE_{max} = qV_{stopping}$ |

Table of work functions

| $\phi_{rodium} = 2.28 eV$ | $\phi_{copper} = 4.7 eV$ | $\phi_{coletum} = 2.9 eV$ | Φ _{pintinum} = 6.35 eV | $\phi_{gold} = 5.1 eV$ |
|----------------------------|---------------------------|----------------------------|---------------------------------|-------------------------|
|----------------------------|---------------------------|----------------------------|---------------------------------|-------------------------|

- What is the energy of each photon for the wavelengths and frequencies shown? Give answers in both Joules and electron-volts.
 - λ=240 nm (ultraviolet) 8.3×10⁻³⁶ J 5.2 eV
 F=104.3 MHz (radio) 6.9×10⁻³⁶ J 4.3×10⁻² eV
 λ=21 cm (microwave) 9.5×10⁻³⁶ J 5.9×10⁻⁴ eV
 λ=650 nm (red light) 3.1×10⁻³⁶ J 1.9 eV

Which of these photons is most likely to produce a current in the photoelectric effect? The 5.2 eV photons are the most likely to produce a current since they have the most energy. Comparing their energy to the work functions shown in the table above they would produce a current for sodium, copper, calcium, and gold, but not for gold. The red light would only produce a current if the target material had a very small work function. Microwaves and radiowaves have much too long a wavelength and too small a photon energy to eject any electrons.

2. Find the wavelength and frequency of photons with the following energies:

| a. | Ephotos=1200 eV (x-ray) | 1.0 nm | 3.0x10 D Hz |
|----|---------------------------------------------|--------|--------------------------|
| b. | Epheno=4.2×10 ⁻²⁸ J (blue light) | 476 nm | 6.3×10 ¹⁴ Hz |
| C. | E _{sterns} =0.51 MeV (gamma) | 2.4 pm | 1.25×10 ³⁰ Hz |

- What is the minimum frequency of light that would produce a photocurrent if the target is made of gold?
 At threshold frequency the energy of the photon is just large enough to remove an electron.
 Hf₀-φ_{ents} → f₀-φ_{ents}/h = 5.1eV/(4.14×10⁻¹⁵ eV s) = 1.23×10²⁵ Hz
- What is the maximum kinetic energy of electrons if radiation with a wavelength of 350 nm shines on sodium?
 KE_{max}=hc/λ f_{midum}=(1240 eV nm)/(350 nm) 2.28 eV = 1.26 eV = 2.02x10⁻¹⁹ J
- 5. When a monochromatic source of light is shining on a copper target a voltage of 3.5 V stops the current.
 - a. What is the maximum speed of the electrons? KE_{max} = qV_s = 3.5 eV = 5.6x10^{±9} J → v=sqrt(2KE/m)=1.1x10⁶ m/s
 - b. What is the wavelength of the light? 3.5 eV = hc/λ ϕ copper $\Rightarrow \lambda=1243$ eV nm / (3.5 eV + 4.7 eV) = 152 nm
- What is the stopping voltage if UV radiation with a wavelength of 100 nm shines on platinum?
 qV_s=hc/λ φ_{platinum} = 12.43 eV 6.35 eV = 6.08 eV so, the stopping voltage is 6.08 V

Photoelectric Effect Problems With Answers

Galen C. Duree, Jr.

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