

Assessment #3 Review

Name: _____

You must show all work for credit.

Due Date: _____ (No late packets will be accepted)

I. Balance Equations*Balance the equations*

- $\text{Mg} + \underline{2} \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + \underline{2} \text{NaBr} \rightarrow \text{MgBr}_2 + \underline{2} \text{Na}$
- $\text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \text{Fe} + \underline{3} \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of Fe_2O_3 in reaction #6? 3

Write the balanced equation for problems # 8 and 9:

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on left, 2 on right.
sub created, or lost

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

2 N, 6 H, 2 on left, 2 on right.
sub created, or lost

12. Using the balanced equation shown in problem #10, list the products of the reaction.

2 NH₃

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

N₂, 3 H₂

II. Types of Reactions*Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



Physical Science Chemistry Test Answers Unit

Zhe Xu



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 - The Rise of Digital Reading Physical Science Chemistry Test Answers Unit
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chemistry Test Answers Unit
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chemistry Test Answers Unit
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chemistry Test Answers Unit

- Personalized Recommendations
 - Physical Science Chemistry Test Answers Unit User Reviews and Ratings
 - Physical Science Chemistry Test Answers Unit and Bestseller Lists
5. Accessing Physical Science Chemistry Test Answers Unit Free and Paid eBooks
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 6. Navigating Physical Science Chemistry Test Answers Unit eBook Formats
 - ePub, PDF, MOBI, and More
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 - Physical Science Chemistry Test Answers Unit Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chemistry Test Answers Unit
 - Highlighting and Note-Taking Physical Science Chemistry Test Answers Unit
 - Interactive Elements Physical Science Chemistry Test Answers Unit
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chemistry Test Answers Unit
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 - Creating a Diverse Reading Collection Physical Science Chemistry Test Answers Unit
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Physical Science Chemistry Test Answers Unit
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Science Chemistry Test Answers Unit

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- Distinguishing Credible Sources

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- Exploring Educational eBooks

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- Integration of Multimedia Elements
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