

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, read about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

octaves	atomic mass	atomic number	nine
elements	properties	Henry Moseley	eight
protons	periodic law	Dmitri Mendeleev	accepted

The table below was developed by John Newlands and is based on a relationship called the law of (1) _____. According to this law, the properties of the elements repeated every (2) _____ elements. Thus, for example, element two and element (3) _____ have similar properties. The law of octaves did not work for all the known elements and was not generally (4) _____.

1	2	3	4	5	6	7
H	Li	G	Bo	C	N	O
8	9	10	11	12	13	14
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to (5) _____. In his table, the elements were arranged according to increasing (6) _____. One important result of this table was that the existence and properties of undiscovered (7) _____ could be predicted.

The element in the modern periodic table are arranged according to increasing (8) _____, as a result of the work of (9) _____. This arrangement is based on number of (10) _____ in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) _____, which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) _____.

Periodic Table Teaching Transparency Answers

**Barbara R. Greenberg,Dianne
Patterson**



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Elements and the Periodic Table, Grades 5 - 8 Abbg, 2013-01-02 Aligned to Common Core State Standards Elements and the Periodic Table present the basics of the Periodic Table in an easy to understand easy to master way It contains fun activities transparency masters quizzes tests rubrics grading sheets and more From basic elements to table organization Elements and the Periodic Table is the essential handbook for middle school science **Periodic Table** Cindy Blobaum, 2005 What do chocolate chip cookies chemistry and logic have in common They are the basis for a unit that lets students become actively engaged in discovering the arrangement of the periodic table This learning activity takes the periodic table out of the static presentation usually associated with textbooks and chemistry courses and interjects an element of discovery The two activities in this unit provide students with information that they have to arrange in organized charts In the process of creating the arrangements students will be involved in problem solving and will gain an appreciation for the scientific process of exploration and verification This dynamic unit meets national science standards in seven teaching and content areas Bring the periodic table to life with this hands on minds on unit Book jacket **Elements and the**

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Speaking and Writing for English Language Learners Dorit Sasson, 2013-10-04 This book provides an overview designed to help educators collaborate more effectively in the areas of content area literacy for the sake of their K 6 ELL students The book weaves the practical and theoretical aspects of collaboration and suggests ways for teachers to form long term partnerships Each chapter extends collaboration in the areas of skill and content based learning so ELL students can achieve necessary proficiency to thrive in content areas classrooms and minimize gaps in instructional learning **Glencoe Science** McGraw-Hill Staff, 2001-06 *Art in Chemistry* Barbara R. Greenberg, Dianne Patterson, 2007-12-30 Integrate chemistry and art with hands on activities and fascinating demonstrations that enable students to see and understand how the science of chemistry is involved in the creation of art Investigate such topics as color integrated with electromagnetic radiation atoms and ions paints integrated with classes of matter specifically solutions three dimensional works of art integrated with organic chemistry photography integrated with chemical equilibrium art forgeries integrated with qualitative analysis and more This is a complete and sequential introduction to General Chemistry and Introductory Art topics In this newly revised edition the author a retired Chemistry teacher gives extensive and in depth new explanations for the experiments and demonstrations as well as expanded safety instructions to insure student safety Grades 7 12 *Holt*

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