

BC Science 8 Workbook Answers

Unit 1: Cells and Systems

Chapter 1 The cell is the basic unit of life.

Section 1.1 Answers

Section 1.1 Summary

Reading Checks

Pages 2–3

1. A microscope helps you focus two objects or details that are close together and makes an object seem larger than it is.
2. The object will appear 16x larger than it is.

Analyzing Information

Characteristics of living things

Page 4

Accept all reasonable answers. Sample answers (can be in any order):

1. responds to its environment—we respond to hunger by eating
2. needs energy—plants need sunlight, water, and carbon dioxide to make food
3. grows—children grow taller, replace cells
4. reproduces—salmon lay eggs
5. gets rid of wastes that build-up in its body—exhaling carbon dioxide

Interpreting Illustrations

The compound light microscope

Page 5

1. light source
2. stage
3. objective lenses
4. eyepiece
5. arm
6. coarse focus knob
7. fine focus knob

Cloze activity

Microscopes

Page 6

1. compound light microscope
2. eyepiece
3. coarse focus knob
4. fine focus knob
5. objective lenses
6. light source

7. resolving power
8. upside down, reversed
9. electron micrograph

Assessment

Observing living things

Page 7

1. B 2. F 3. C 4. G 5. E 6. A 7. D 8. B 9. D 10. A 11. C 12. A 13. B

Section 1.2 Answers

Section 1.2 Summary

Reading Checks

Pages 8–9

1. Prokaryotic cells do not have organelles with a membrane around them.
2. cell wall, chloroplasts

Interpreting Illustrations

Parts of cells

Page 10

- A. cell membrane
- B. nucleus
- C. mitochondria
- D. vacuoles
- E. cytoplasm
- F. mitochondria
- G. cytoplasm
- H. nucleus
- I. chloroplast
- J. vacuole
- K. cell wall
- L. cell membrane

Cloze Activity

Inside a cell Page 11

1. organelle
2. cell membrane
3. cytoplasm
4. nucleus
5. mitochondria
6. vacuoles
7. cell wall
8. chloroplasts
9. eukaryotic

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Katrin Zwirgmaier



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web other articles where exothermic reaction is discussed alkali metal reactions with nonmetals the reactions are highly exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 via and via of the periodic table sulfides can be formed by the direct reaction of the alkali metals with

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable

measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol⁻¹

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic

eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹ eg 2 HNO₃(aq) + KOH(aq) → KNO₃(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example HCl(aq) + NaOH(aq) → NaCl(aq) + H₂O(aq) you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts

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