

**CHAPTER 6 REVIEW***Chemical Bonding***MIXED REVIEW**

**SHORT ANSWER** Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ <b>lattice energy</b> _____	<b>a.</b> ionic bond
_____ <b>bond energy</b> _____	<b>b.</b> covalent bond
_____ <b>enthalpy of vaporization</b> _____	<b>c.</b> metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ <b>ionic</b> _____	<b>a.</b> H—F	_____ <b>nonpolar covalent</b> _____	<b>d.</b> H—H
_____ <b>ionic</b> _____	<b>b.</b> Na—Cl	_____ <b>polar covalent</b> _____	<b>e.</b> H—C
_____ <b>polar covalent</b> _____	<b>c.</b> H—O	_____ <b>polar covalent</b> _____	<b>f.</b> H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen  
atom and the unshared electron pair of a strongly electronegative atom such as O,  
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,  
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally  
intermolecular, while ionic and covalent bonds occur between ions or atoms  
respectively.

4. H<sub>2</sub>S and H<sub>2</sub>O have similar structures and their central atoms belong to the same group. Yet H<sub>2</sub>S is a gas at room temperature and H<sub>2</sub>O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.  
Increased polarity in H<sub>2</sub>O bonds means a stronger intermolecular attraction, making  
water a liquid at room temperature. Hydrogen bonding exists between water  
molecules, but not between hydrogen sulfide molecules.

# Covalent Bonding Chapter 8 Worksheet Answers

## Alitaoore

**M Carnoy**



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### **Table of Contents Covalent Bonding Chapter 8 Worksheet Answers Alitaoore**

1. Understanding the eBook Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - The Rise of Digital Reading Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Personalized Recommendations

- Covalent Bonding Chapter 8 Worksheet Answers Alitaoore User Reviews and Ratings
- Covalent Bonding Chapter 8 Worksheet Answers Alitaoore and Bestseller Lists
- 5. Accessing Covalent Bonding Chapter 8 Worksheet Answers Alitaoore Free and Paid eBooks
  - Covalent Bonding Chapter 8 Worksheet Answers Alitaoore Public Domain eBooks
  - Covalent Bonding Chapter 8 Worksheet Answers Alitaoore eBook Subscription Services
  - Covalent Bonding Chapter 8 Worksheet Answers Alitaoore Budget-Friendly Options
- 6. Navigating Covalent Bonding Chapter 8 Worksheet Answers Alitaoore eBook Formats
  - ePub, PDF, MOBI, and More
  - Covalent Bonding Chapter 8 Worksheet Answers Alitaoore Compatibility with Devices
  - Covalent Bonding Chapter 8 Worksheet Answers Alitaoore Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Highlighting and Note-Taking Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Interactive Elements Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
- 8. Staying Engaged with Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
- 9. Balancing eBooks and Physical Books Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Setting Reading Goals Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonding Chapter 8 Worksheet Answers Alitaoore
  - Fact-Checking eBook Content of Covalent Bonding Chapter 8 Worksheet Answers Alitaoore

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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