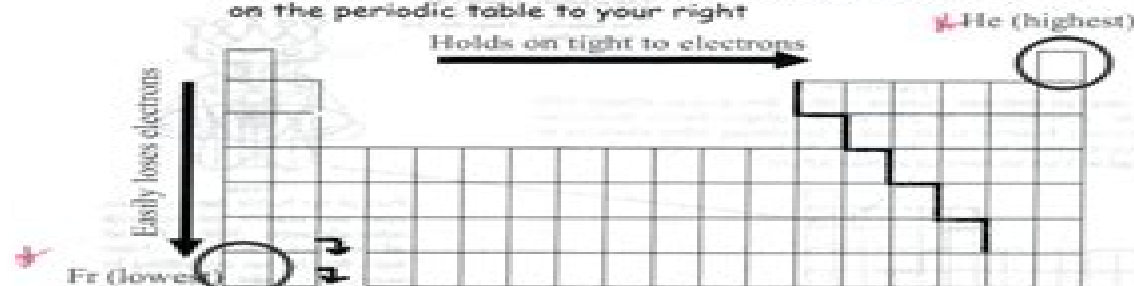


b. Why do elements get larger as you go down a group? Additional energy levels are added to the electron cloud as you go down a group.

### Ionization Energy:

1. In simple terms, explain ionization energy. The energy required to remove an electron from a neutral atom.
2. Using the information in figure 5-15 (Pg 143) in your book, write the ionization energy for alkali metals & period #2 on the periodic table to your right



3. On the arrows on the diagram, indicate the trend for ionization energy as you move down a column and across a row. (Options: easily loses an electron OR holds on tight to electrons)
4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest ionization energy.

5. Rank the following elements from highest to lowest ionization energy:

a) Be, Ba, Mg = Be, Mg, Ba

b) Li, F, B = F, B, Li

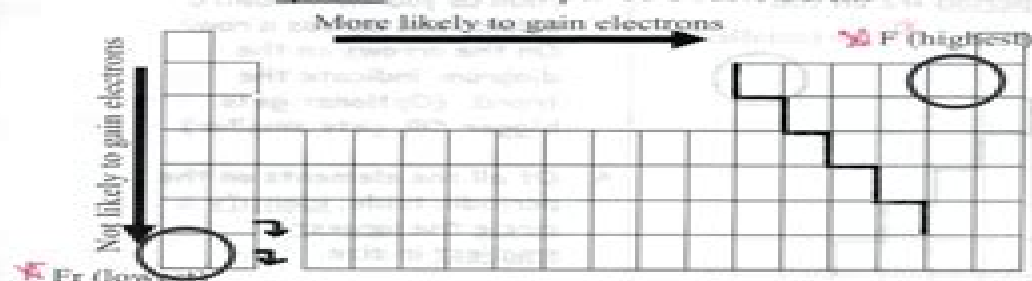
6. Why does it take less energy to remove an electron from a large atom compared to a small atom? Their electrons are closer to the nucleus, so the protons have a larger "pull" on them.

*It takes less energy to remove an electron from a large atom because*

Electronegativity: the effective nuclear charge decreases and electrons are not held as tightly since

1. Explain Electronegativity: The ability of an atom to get electrons from other atoms.

2. Using the information in figure 5-20 (pg 151) in your book, write the electronegativity for the Halogens and period #2 on the periodic table below.



3. On the arrows on the diagram, indicate the trend for electronegativity as you move down a column and across a row. (Options: more likely to gain electrons, not likely to gain electrons)

4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest electronegativity.

5. Rank the following elements from most to least electronegative element:

a) Be, Ba, Mg = Be, Mg, Ba

b) Li, F, B = F, B, Li

6. Why are small atoms better at pulling electrons off of larger atoms than vice versa?

Their electrons are closer to the nucleus, so the protons have a larger "pull" on the electrons in other atoms.

*not detailed enough*  
*The increased number of energy level(s) increases the distance over which nucleus must pull and therefore reduces the attraction for electrons* } *AP answer*

# Cracking The Periodic Table Code Pogil Answers

**Katrin Zwirgmaier**



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