

Honors Chemistry I
Stoichiometry Practice Test

Note: This practice test has more questions on it than you will see on the actual test. I have tried to include a large variety of questions so that you will be prepared for Thursday's test!

I. A student dissolves 3.01 grams of solid copper metal in an excess of 6.00 M nitric acid, producing gaseous nitric oxide (NO), cupric nitrate, and water. Once the gas was removed, the blue solution was heated to drive off all of the water and 8.70 grams of a blue precipitate were recovered. The reaction was conducted at STP conditions. Answer the following questions based on this information.

A. Write the balanced equation for this reaction. [Hint: the sum total of all coefficients adds up to 20.]



B. What volume of nitric oxide should have been formed?

$$3.01 \text{ g Cu} \times (1 \text{ mol}/63.55 \text{ g}) \times (2 \text{ NO}/3 \text{ Cu}) \times (22.41 \text{ L}/1 \text{ mol}) = .707 \text{ L}$$

C. What is the mass of water produced?

$$3.01 \times (1 \text{ mol}/63.55 \text{ g}) \times (4\text{H}_2\text{O}/3\text{Cu}) \times 18.016 \text{ g}/1 \text{ mol}) = 1.14 \text{ g}$$

D. How many moles of nitric acid were needed initially?

$$3.01 \text{ g Cu} \times (1 \text{ mol}/63.55 \text{ g}) \times (8\text{HNO}_3/3\text{Cu}) = 0.126 \text{ mol}$$

E. How many grams of cupric nitrate should have been produced, assuming 100% yield?

$$3.01 \text{ g} (1 \text{ mol}/63.55 \text{ g}) \times (3\text{Cu}(\text{NO}_3)_2/3\text{Cu}) \times (187.57 \text{ g}/1 \text{ mol}) = 8.88 \text{ g}$$

F. What is the percent yield for this reaction?

$$8.70/8.88 \times 100 = 98.0 \%$$

G. If the student wanted to add exactly the stoichiometric amount of acid needed to dissolve the copper metal, given that the concentration of the acid was 6.00 M, what volume of acid would have been needed?

$$6.00 \text{ M} = .126 / x \quad x = 0.021$$

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