

Identifying unknown using stoichiometry

Experiment 2

Experiment performed on 2/22/2021

Amjad Samra

Lab partners: Antonio Alderete, Tiana Marie and
Allie Howard

Chemistry 1215L section 12

Hypothesis:

The only difference between sodium carbonate and sodium bicarbonate is the amount of CO₂ produced during the reaction. Since the CO₂ production is the only component of the reaction that would cause a significant change in pressure, the pressure measurements of the reactions of both unknown chemicals with HCl can provide a direct link to the stoichiometry of CO₂.

Chemicals and hazard:

Chemical name	Molecular formula	Molar mass	Amount used
Sodium chloride	NaCl	58.44 g/mL	'product'
Sodium carbonate	Na ₂ CO ₃	105.989 g/mL	25 mL
Hydrochloric acid	HCl	36.458 g/mL	50 mL
water	H ₂ O	18.01528 g/mL	'product'
Carbon dioxide	CO ₂	44.01 g/mL	'product'
Sodium hydrogen carbonate	NaHCO ₃	88.007 g/mL	25 mL

Identifying Unknown Solutions Chemistry Lab Report

Georgios Tsaparis



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Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures **Inquiry-based Experiments in Chemistry** Valerie Ludwig Lechtanski, 2000 Inquiry Based Experiments in Chemistry is an alternative to those cookbook style lab manuals providing a more accurate and realistic experience of scientific investigation and thought for the high school chemistry or physical science student *Laboratory Course Chemistry* Mr. Rohit Manglik, 2024-03-15 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels **Practical Chemistry** Kasaye Bahiru Tola, Girma Salale Geleta, 2025-03-03 This laboratory manual offers a broad introduction to practical instrumental analysis and quantitative analytical chemistry The practical activities include experiments to determine the quantity of analytes Analytical techniques covered in the book are turbidimetry atomic absorption spectrometry flame emission spectrometry refractometry infrared spectroscopy fluorometry and UV visible spectrophotometry Practical Chemistry in 3 Volumes Volume 1 Practical Chemistry Instrumental Analysis ISBN 978 3 11 157504 9 Volume 2 Practical Chemistry

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between students and their lab instructors to be two defining factors for lab course success was used as a framework

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