

QUALITATIVE ANALYSIS of GROUP III CATIONS

The four group III cations are Cr^{3+} , Al^{3+} , Fe^{3+} , and Ni^{2+} . The first step in analysis involves separating the ions into two subgroups by treating the solution with NaOH and NaOCl . The hypochlorite ion oxidizes Cr(III) to a higher, more stable oxidation state (namely Cr(VI)) which is soluble:



In addition, Al^{3+} forms a soluble hydroxo-complex ion in the presence of excess hydroxide:



In contrast, Ni^{2+} and Fe^{3+} do not readily form hydroxo-complexes and are not oxidized by hypochlorite. They form insoluble hydroxides under these conditions:



To separate aluminum from chromium, the solution containing CrO_4^{2-} and Al(OH)_4^{-} is acidified to destroy the hydroxo-complex:



Treatment with aqueous ammonia gives a gelatinous **white** precipitate of aluminum hydroxide. The concentration of hydroxide in ammonia is too low to form the hydroxo-complex:



The chromate ion remains in solution. It can be tested and confirmed by precipitation as yellow BaCrO_4 :



The BaCrO_4 precipitate dissolves in acid. The solution is then treated with H_2O_2 to produce a **deep blue** color due to the presence of a peroxo-compound, probably CrO_5 :



The mixed precipitate of Ni(OH)_2 and Fe(OH)_3 can be dissolved by adding a strong acid. The Ni^{2+} and Fe^{3+} ions can be separated by adding ammonia. Ni^{2+} is converted to the deep-blue complex $\text{Ni(NH}_3)_6^{2+}$ which stays in solution. The Fe^{3+} ion does not readily form a complex and re-precipitates as Fe(OH)_3 :



Confirm the presence of Ni^{2+} by adding dimethylglyoxime (DMG), $\text{C}_4\text{H}_8\text{N}_2\text{O}_2$, to give a **deep rose** precipitate:



Confirm the presence of Fe^{3+} by dissolving Fe(OH)_3 in a strong acid and adding KSCN to form a **blood-red** FeSCN^{2+} complex ion:



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