

4. a. (i) When conc. = 0.4 mol.l^{-1} , rate = 0.2 s^{-1}
 so time = $\frac{1}{\text{rate}} = \frac{1}{0.2} = 5 \text{ s}$
- (ii) As the concentration increases, the number of collisions increases, so I increasing the rate of the reaction.
- b.(i) Select a rate from the graph and double it. Find the temp. associated with these rates and the difference gives the temp.rise.
- eg Rate 1 = 2 s^{-1} Temp. = 40°C
 Rate 2 = 4 s^{-1} Temp. = 55°C
 so temp. rise = $55 - 40$ = 15°C
- (ii) As the temperature increases, the particles gain more energy. This increases the number of successful collisions.
5. a. Rate during the first 10 seconds = $5.5 \times 10^{-3} \text{g/s}$
 b. Rate between 20 and 40 seconds = $1.1 \times 10^{-3} \text{g/s}$
6. a. Rate during the first 20 seconds = $1.5 \times 10^{-3} \text{mol/l/s}$
 b. Rate between 30 and 70 seconds = $3.75 \times 10^{-3} \text{mol/l/s}$
7. a. The amount of sulfur needed to obscure the cross is the same for each experiment, so the rate of reaction is proportional to $1/t$.
 b. approx. 150 seconds
 c. approx. 0.19mol l^{-1}
8. a. The readings show a decrease in mass because the gas carbon dioxide is given off.

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