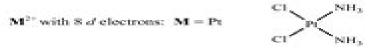
b. [M(H₂O)_o]³⁺ M is second row transition metal, LFSE = -2.4 Δ_o. This can be achieved with a low spin d⁶ configuration. M³⁺ with 6 d electrons: M = Rh



- c. [MCl₄] M is first row transition metal, 5 unpaired electrons.

 M^{3*} with 5 d electrons: Fe
- d. MCl₂(NH₃)₂ M is third row d⁶ transition metal; two M-Cl stretching bands in IR. Second and third row d⁶ complexes are often square planar. The presence of two M-Cl stretching bands implies cis geometry of the chloro ligands.



- Angular overlap calculations for d* and d* ions show no energy difference between D_{th} and O_b when exclusively considering a interactions. Both d* geometries have energies of -3c_n; both d* geometries have energies of -6c_n. In general, stability constants decrease as more ligands are added; the sequence for nickel is the common one. The huge drop in stability constant between the second and third ethylenediamine on Cu²⁺ is a result of the d* Jahn-Teller effect. The first two en ligands add in a square-planar geometry, with water molecules in the axial positions, and the coordination of these two mondentate ligands allows for the Jahn-Teller distortion. Adding a third en ligand requires a geometry change and with a preference for uniform M—N bond distances towards the six mitrogen atoms. This is counter to the Jahn-Teller distortion, and the third addition is much less favorable than the first two.
- 10.5 $[M(H_2O)_6]^{3^\circ}$ M is first row transition metal, $\mu = 3.9$ BM. The magnetic moment implies 3 unpaired electrons, which would give rise to $\mu = \sqrt{3(3+5)} = 3.9$ BM. There are two possibilities, d^3 and d^3 :

$$\mathbf{M}^{2^{+}}, d^{3} \colon \mathbf{M} = \mathbf{V} \qquad \qquad \mathbf{M}^{2^{+}}, d^{7} \colon \mathbf{M} = \mathbf{Co}$$

$$\underline{\bullet} \quad \underline{\bullet} \quad \underline{\bullet} \quad \underline{\bullet} \quad \underline{\bullet}$$

$$\mathbf{10.6} \quad \mathbf{a.} \qquad [\mathbf{Cr}(\mathbf{H}_{2}\mathbf{O})_{\mathbb{R}}]^{2^{+}} \quad n = 4$$

$$\mu = \sqrt{4(6)} = 4.9 \ \mu_{\mathbb{B}}$$

$$\mathbf{b.} \qquad [\mathbf{Cr}(\mathbf{CN})_{\mathbb{S}}]^{4^{+}} \quad n = 2$$

$$\mu = \sqrt{2(4)} = 2.8 \ \mu_{\mathbb{B}}$$

$$\mathbf{c.} \qquad [\mathbf{FeCl}_{4}] \qquad n = 5$$

$$\mu = \sqrt{5(7)} = 5.9 \ \mu_{\mathbb{B}}$$

$$\mathbf{d.} \qquad [\mathbf{Fe}(\mathbf{CN})_{\mathbb{S}}]^{3^{+}} \quad n = 1$$

$$\mathbf{M}^{2^{+}}, d^{7} \colon \mathbf{M} = \mathbf{Co}$$

$$\mu = \sqrt{1(3)} = 1.7 \ \mu_{\mathbb{B}}$$

$$\mathbf{f.} \qquad [\mathbf{Cu}(\mathbf{en})_{2}(\mathbf{H}_{2}\mathbf{O})_{2}]^{2^{+}} \quad n = 1$$

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Gary L. Miessler, Donald Arthur Tarr

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