

FORMULAS AND OXIDATION NUMBERS

A chemical formula is a combination of symbols and numerical subscripts that represents the composition of a compound. The symbols indicate which elements are present and the numerical subscripts indicate the relative proportion of each element in the compound. These proportions can be predicted using the oxidation numbers of the elements and the charges of polyatomic ions. When atoms acquire a charge they are called ions. Ions consisting of more than one atom are polyatomic ions. Its oxidation number represents the apparent charge on an atom.

It is important that all scientists use the same system for writing chemical formulas. This helps to ensure clear and consistent transmission of information. Therefore, the following rules should be used for writing chemical formulas.

1. In a neutral compound the sum of the oxidation numbers of the elements and the charges on polyatomic ions in that compound must equal zero.
2. One positive (+) charge will neutralize one negative (-) charge.
3. Atoms with positive oxidation numbers or ions with positive charges are written first.
4. When the relative proportion of an element in a compound is greater than one, the symbol for that element must be followed by a numerical subscript indicating its relative proportion, as in MgCl_2 .
5. When the relative proportion of a polyatomic ion in a compound is greater than one, the symbol for the polyatomic ion must be enclosed by parentheses, followed by the correct numerical subscript, as in $\text{Al}_2(\text{SO}_4)_3$.

In this experiment you will use cut-out models of ions to form neutral compounds. The correct chemical formula and name for each compound will be determined by balancing oxidation numbers and charges.

OBJECTIVES

In this experiment, you will:

- cut out models of ions
- match the necessary number of ions until the positive and negative oxidation numbers equal 0
- predict the correct formulas for the compounds listed.

EQUIPMENT

scissors
pencil and paper
sheet of ion models
glue stick

PROCEDURE

1. Prepare a data table as directed in the Analysis.
2. Cut out each of the "ion" squares from the sheet provided by your teacher.
3. Construct formulas for the following combinations of substances. For example, the formula for a compound containing magnesium and chlorine may be determined in the following way: Place the Mg^{2+} ion on a piece of paper. Place enough Cl^- ions alongside the Mg^{2+} ion to balance the charges. (Positive charges must equal negative charges.)

4. Predict formulas for four additional compounds using the ions listed in this experiment. (see data table for requirements). Use references to determine if your compounds exist.

Combining Substances:

aluminum and bromine
sodium and oxygen
magnesium and sulfur
aluminum and nitrate ion
potassium and sulfate ion
ammonia and chlorine
ammonium ion and sulfur
aluminum and oxygen
magnesium and sulfate ion
sodium and phosphate ion

Formulas And Oxidation Numbers Lab Answers

M Mosston



Formulas And Oxidation Numbers Lab Answers:

Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16
Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

Laboratory Manual Jo A. Beran, 1990
Environmental Sampling and Analysis Maria Csuros, 1997-03-24 This manual covers the latest laboratory techniques state of the art instrumentation laboratory safety and quality assurance and quality control requirements In addition to complete coverage of laboratory techniques it also provides an introduction to the inorganic nonmetallic constituents in environmental samples their chemistry and their control by regulations and standards Environmental Sampling and Analysis Laboratory Manual is perfect for college and graduate students learning laboratory practices as well as consultants and regulators who make evaluations and quality control decisions Anyone performing laboratory procedures in an environmental lab will appreciate this unique and valuable text

Nuclear Science Abstracts, 1975 **Technical Publications Announcements with Indexes** United States. National Aeronautics and Space Administration, 1962 **Energy Research Abstracts**, 1993 *ERDA Energy Research Abstracts*, 1983 *Illustrated Guide to Home Chemistry Experiments* Robert Bruce Thompson, 2008-04-29 Provides information on setting up an in home chemistry lab covers the basics of chemistry and offers a variety of experiments **Scientific and Technical Aerospace Reports**, 1994-03 College Credit

Recommendations ,1996 *Studies on Life at the Energetic Edge – from Laboratory Experiments to Field-Based Investigations, Volume II* Mark Alexander Lever,Bo Barker Jorgensen,Jan Amend,Victoria Orphan,Tori Hoehler,2024-01-16 In collaboration with Microenergy 2022 The 4th International Workshop on Microbial Life under Extreme Energy Limitation we are proud to launch Volume II of Studies on Life at the Energetic Edge from Laboratory Experiments to Field Based Investigations This workshop focuses on the energy controls on microbial life and the exploration of the biological demand for energy Genetic adaptations and phenotypic traits that enable microorganisms to tolerate long periods of energy limitation have attracted broad scientific interest in recent years Laboratory based cultivation experiments have shown that the potential to survive weeks to months in the absence of energy inputs occurs across a phylogenetically wide range of microbes Studies on natural environments have shown that energy limitation is pervasive across most habitats on Earth from highly metabolically active surface habitats to subsurface environments that have been cut off from new energy inputs for thousands of years Yet much remains to be learned about the evolutionary adaptations and life history traits that enable microorganisms to live under low energy conditions Similarly the spectrum of energy sources and metabolisms that enable and support life on Earth and potentially elsewhere in the Universe is far from constrained *Chemistry* McGraw-Hill Staff,2001-07 **ERDA Energy Research Abstracts** United States. Energy Research and Development Administration,1977 **Electrochemistry** Susan Marie Szegedi,2000 **History of Met Lab Section C-I** Glenn Theodore Seaborg,1977 Study Guide for Chemical Principles Thomas Elliott Taylor,1979 Essentials of Chemistry Ralph A. Burns,John William Hill,1995 The second edition of this chemistry textbook that uses practice examples and applications relating chemistry to our lives and the environment Fundamentals of Chemistry Ralph A. Burns,1995 **Laboratory Manual to Accompany Introductory Chemistry** Uno Kask,Otto W. Nitz,Martin G. Ondrus,Tammy J. Melton,1993-07 This fifth edition of this laboratory manual emphasizes safety in the lab and discusses equipment requirements in the apparatus section at the beginning of each experiment It also features a revised art programme and explains the rational for each experiment

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