

Determining the Vapor Pressure of a Gas by P-T Extrapolation

OBJECTIVES

It is proposed that experiments be conducted that involve the volume of one mole of hydrogen gas at standard temperature and pressure (STP). Hydrogen will be generated, the volume collected above some of hydrogen will be known, and the hydrogen gas manometrically added until some known. The volume of hydrogen added by water displacement will be calculated and corrected for difference in temperature and pressure in order to compare the actual volume of hydrogen at STP.

Background Information

Avogadro's Law states that equal volumes of gases contain equal numbers of molecules under the same conditions of temperature and pressure. In laboratory situations, this will give us hydrogen containing the same number of molecules with volume the same volume if the temperature and pressure are kept constant. The volume occupied by one mole of gas is called the molar volume. In this experiment, the molar volume of hydrogen gas at standard temperature and pressure (STP) equal to 22.4 L and is used to compare and contrast.

The volume of hydrogen added with hydrochloric acid & magnesium is provided, a correction factor of generating hydrogen in the lab.

Figure 1: Hydrogen Gas Collection Setup

Equation 1

If the reaction occurred in a well sealed hydrogen gas, the volume of hydrogen gas collected will depend on the amount of material magnesium as well as the pressure and temperature. The volume of hydrogen gas is calculated if the volume occupied by a known number of moles of hydrogen is known. Since the volume will be measured under laboratory conditions of temperature and pressure, the measured volume for comparison STP conditions have a correction factor known.

The relationship among the four gas variables: pressure (P), volume (V), temperature (T), and the number of moles (n), is expressed in the ideal gas law, Equation 2, where R is a constant called the universal gas constant.

$$PV = nRT$$

Equation 2

If the ideal gas law under the Equation 2, the number of moles of gas is known, the volume of gas, the gas has used in collecting the volume (V) and a gas at STP (T) and P is known, the volume of hydrogen gas under any other gas of laboratory conditions (T) and P, the relationship that ideal gas law in the number of gas law, substituted the temperature and pressure is substituted in constant factor to give the volume, compare with.

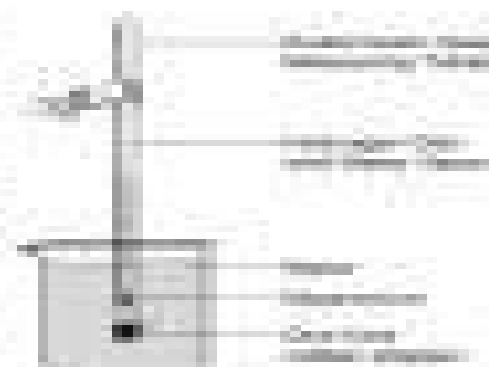
$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Equation 3

Hydrogen gas will be collected by the displacement of water in an inverted gas collecting tube held at a constant level using the apparatus shown in Figure 1. The total pressure of the gas inside will be the equal to the atmospheric gas pressure. However, the gas in the cylinder contains the gas hydrogen. The gas will be a water-vapor vapor since the composition of the atmosphere is not 100% hydrogen. In fact, according to Dalton's Law, the total pressure of the gas will be equal to the partial pressure of hydrogen plus the partial pressure of water vapor (Equation 4). The vapor pressure of water depends only on the temperature (see Table 1).

$$P_{\text{total}} = P_{\text{H}_2} + P_{\text{H}_2\text{O}}$$

Equation 4



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