

Determination of Iron in Ore by Redox Titration

Oxidation-reduction reactions are suitable for titration reactions if they are rapid and quantitative. The end point can be seen by indicators or by instrumental methods such as potentiometry.

Potentiometric Titration of Ferrous Ion

The reaction of ferrous ion with ceric ion meets the above requirements.



The course of the titration may be followed potentiometrically with a platinum electrode. The following cell may be used:



The potential of the Pt electrode must obey the Nernst equation since the half reactions between Fe^{3+} and Fe^{2+} and between Ce^{4+} and Ce^{3+} are nearly reversible.

$$E_{\text{Pt}} = E^{\circ}(\text{Fe}^{3+}, \text{Fe}^{2+}) + 0.05916 \log \left(\frac{[\text{Fe}^{3+}]}{[\text{Fe}^{2+}]} \right) = E^{\circ}(\text{Ce}^{4+}, \text{Ce}^{3+}) + 0.05916 \log \left(\frac{[\text{Ce}^{4+}]}{[\text{Ce}^{3+}]} \right) \quad (2)$$

Equation 2 states that the Pt electrode can have only one potential, and the ratios of $[\text{Fe}^{3+}]/[\text{Fe}^{2+}]$ and $[\text{Ce}^{4+}]/[\text{Ce}^{3+}]$ must be such that the Nernst equation is satisfied for both half reactions. Reaction 1 will occur spontaneously to such an extent that equation 2 will be true. The ratios $[\text{Fe}^{3+}]/[\text{Fe}^{2+}]$ and $[\text{Ce}^{4+}]/[\text{Ce}^{3+}]$ change greatly at the equivalence point. The equivalence point is determined from the break in the titration curve.

Indicator Titration of Ferrous Ion

Specific indicators and true oxidation-reduction indicators are available for redox titrations. Specific indicators are substances that react with one or more of the participants in a titration to produce a color change. A classic example is starch that forms a dark blue complex with I_2 . In redox titrations with I_2 as the titrant, blue color develops to mark the end point of the reaction when I_2 begins to accumulate in the solution. Thiocyanate is an indicator that forms a red complex with Fe^{3+} but not with Fe^{2+} , and thus can be used to indicate the formation or disappearance of Fe^{3+} . Sometimes a redox reagent itself is deeply colored and can be used to indicate the end point. For example, MnO_4^- is a purple ion often used as an oxidizing agent. The first trace of purple color indicates that the MnO_4^- is no longer being reduced by the titration reaction.

A true oxidation-reduction indicator is a substance that has different colors when in different oxidation states.



Determination Of Iron In Ore By Redox Titration Chemistry

Leicester Forsyth Hamilton



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Analytical Chemistry of Complex Matrices, 2013-03-08 Analytical Chemistry Gary D. Christian, Purnendu K. Dasgupta, Kevin A. Schug, 2013-10-07 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry , **Comprehensive Engineering Chemistry** Devender Singh, Satish Kumar Vats, 2013-12-30 This book is designed to meet the requirement of the students of B Tech and B E students The book discusses in detail the following topics Thermodynamics Phase Rule Water and its Treatment Corrosion and its Prevention Lubrication and Lubricants Polymer and Polymerization and Analytical Methods The book is suitably illustrated with diagrams and a number of solved numerical examples from different universities are included to make the text more exhaustive and understandable Practical part is also appended at the end of the book **Engineering Chemistry with Laboratory Experiments** Kaurav M. S., 2011 **Chemical Equilibrium and Analysis** Richard W. Ramette, 1981 **Quantitative Chemical Analysis** Na Li, John J Hefferren, Ke'an Li, 2013-04-26 This book covers both fundamental and practical aspects of chemical analysis Data Process and Analysis Chemical Equilibria and Volumetric titrations Gravimetry Spectrophotometry Sample Preparation and Separation Methods in Quantitative Analysis It was written with the rich tradition of teaching at Peking University College of Chemistry and edited by an American professor who was personally sensitive to the needs of students learning science from traditional chemistry textbooks written in English Many examples and illustrative problems in this text have been taken from previous textbooks by the Peking University Team Teaching Program The book can be used as a starter in analytical chemistry which is fundamental and the base upon which chemistry is built Traditional chapters of initial learning in analytical chemistry are included such as volumetric gravimetric and separation methods the book also includes key chapters on problem solving relating to recent progress in analytical chemistry **R - Sm, Berichtsjahre 1932 bis 1962, Lieferung 5** Johann Christian Poggendorff, 2022-09-06 Keine ausf hrliche Beschreibung f r R Sm Berichtsjahre 1932 bis 1962 Lieferung 5 verf gbar SME Mineral Processing and Extractive Metallurgy Handbook Courtney A. Young, 2019-02-01 This landmark publication distills the body of knowledge that characterizes mineral processing and extractive metallurgy as disciplinary fields It will inspire and inform current and future generations of minerals and metallurgy professionals Mineral processing and extractive metallurgy are atypical disciplines requiring a combination of knowledge experience and art Investing in this trove of valuable information is a must for all those involved in the industry students engineers mill managers and operators More

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Hamilton, 1958. **A Brief Introduction to Quantitative Chemical Analysis** Robert Blanchard Fischer, Dennis G.

Peters, 1969. Nature Of Chemistry Volume - 2 Aditya Tripathi, AP Chemistry with Online Tests Neil D.

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Table of Contents Determination Of Iron In Ore By Redox Titration Chemistry

1. Understanding the eBook Determination Of Iron In Ore By Redox Titration Chemistry
 - The Rise of Digital Reading Determination Of Iron In Ore By Redox Titration Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Determination Of Iron In Ore By Redox Titration Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Determination Of Iron In Ore By Redox Titration Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Determination Of Iron In Ore By Redox Titration Chemistry
 - Personalized Recommendations
 - Determination Of Iron In Ore By Redox Titration Chemistry User Reviews and Ratings
 - Determination Of Iron In Ore By Redox Titration Chemistry and Bestseller Lists

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 - Determination Of Iron In Ore By Redox Titration Chemistry Public Domain eBooks
 - Determination Of Iron In Ore By Redox Titration Chemistry eBook Subscription Services
 - Determination Of Iron In Ore By Redox Titration Chemistry Budget-Friendly Options
6. Navigating Determination Of Iron In Ore By Redox Titration Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Determination Of Iron In Ore By Redox Titration Chemistry Compatibility with Devices
 - Determination Of Iron In Ore By Redox Titration Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Determination Of Iron In Ore By Redox Titration Chemistry
 - Highlighting and Note-Taking Determination Of Iron In Ore By Redox Titration Chemistry
 - Interactive Elements Determination Of Iron In Ore By Redox Titration Chemistry
8. Staying Engaged with Determination Of Iron In Ore By Redox Titration Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Determination Of Iron In Ore By Redox Titration Chemistry
9. Balancing eBooks and Physical Books Determination Of Iron In Ore By Redox Titration Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Determination Of Iron In Ore By Redox Titration Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Determination Of Iron In Ore By Redox Titration Chemistry
 - Setting Reading Goals Determination Of Iron In Ore By Redox Titration Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Determination Of Iron In Ore By Redox Titration Chemistry
 - Fact-Checking eBook Content of Determination Of Iron In Ore By Redox Titration Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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