

## DETERMINING AD VALOREM FEE LEVEL

[MCL 207.8011 (a)(2)]

The law which establishes ad valorem fees also provides an original fee level and three levels of depreciation which are used as the vehicle ages. Normally, the "0" level is used for new or current model year vehicles. The "1" level is used for one-year-old vehicles. The "2" level is used for two-year-old vehicles, and the "3" level is used for three-year-old vehicles. Once the license plate fee reaches the "3" level, it no longer changes as the vehicle continues to age.

There is one important variation in determining the plate fee level. If a vehicle owner transfers their license plate to a vehicle at the time of purchase, then only the \$2 transfer fee is collected. When the vehicle owner renames the license plate at a later date, they pay the fee level that would have been due on the date the plate was transferred. Thus, the level of depreciation may be delayed.

For a vehicle purchased as new, fees must be paid at the "0" level before fees are collected at the "1" level, and must progress to the "3" level. A transfer of vehicle ownership breaks this formula. (See used vehicle fee levels on the next page.) The Department's computer always keeps track of the next level due.

Examples of what happens with the levels of used vehicles:

- 1) An owner buys a new 2017 model in November, 2016. Her birthday is in June so the plate is issued for 6 months at the "0" level. When she renames in June, 2017, the owner pays the "1" level fee. The computer keeps track of the next level due. The method of determining the level changes only if ownership is transferred. If she sells the car in July, the new owner pays the "0" level. (See Used Vehicle instructions on the next page.)
- 2) An owner buys a new 2016 model in December, 2016. His birthday is in April so the plate is issued for 17 months at the "0" level. When he renames in April, 2018, the owner pays the "1" level, not the "2" level.
- 3) An owner buys a new 2016 model in early 2017. He transfers a plate expiring 02-12-2018. When the plate is renamed in 2018, the "0" level fee is paid even though the vehicle is already two model years old. (The owner never paid the "0" or "1" level fees on the new vehicle.)

-continued-

# Determining Ad Valorem Fee Level Michigan Som

**Yan Bai**



## **Determining Ad Valorem Fee Level Michigan Som:**

The Michigan Assessor ,1981      **Michigan Compiled Laws Annotated** Michigan,1967      **ITC Publication** United States International Trade Commission,1978      *Michigan Compiled Laws, Annotated* Michigan. Laws, Statutes, Etc,2005  
Michigan Investor ,1943      Appraisal of Railroad and Other Public Utility Property for Ad Valorem Tax Purposes  
National Association of Tax Administrators (U.S.). Committee on Unit Valuation,1954      **A Survey of the Tax Situation in Michigan** Harcourt L. Caverly,1933      *Sugar and Sweetener Report* ,1979      **The Japan Daily Mail** ,1889  
Congressional Record United States. Congress,1948      Standard & Poor's Creditweek ,1988-07      **Issues and Options for U.S.-Japan Trade Policies** Robert Mitchell Stern,2002 Addresses the central negotiating issues involving the trade policies and relations between the United States and Japan      **Foreign Commerce Weekly, Volume 62, Number 1** ,1959      *Engineering* ,1867      **National Stockman and Farmer** ,1889      *The Lawyers Reports Annotated* ,1915  
Journal of the House of Representatives of the United States United States. Congress. House,1993 Some vols include supplemental journals of such proceedings of the sessions as during the time they were depending were ordered to be kept secret and respecting which the injunction of secrecy was afterwards taken off by the order of the House      Poor's Cumulative Service ,1929      **American Economist** ,1917      **State & Local Taxes Weekly** ,1991 Releases consist of report bulletins and legislative bulletins

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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nitration of bromobenzene lab report improved essays 731 words 3 pages open document essay sample check writing quality show more this experiment was performed in order to demonstrate the principles of nitration through electrophilic aromatic substitution it also examined the effects of bromine on an aromatic ring ortho para vs meso directing

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then actual is your actual findings yield actual theoretical  $\times 100$  what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the

nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0 6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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lab report 15 nitration of bromobenzene brynn abbott amp katie muerer bea0006 uah ch 336 02 september 21 2022 abstract in this experiment the electrophilic aromatic substitution mechanism was used to obtain 4 nitrobromobenzene and 2 nitrobromobenzene recrystallization was used to purify 4 nitrobromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no2 which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

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chem 2261 lab 5 nitration of bromobenzene introduction an electrophilic aromatic substitution reaction is an important process that allows for intermediate chemical structures to be used as precursors to pharmaceuticals and other industrially produced chemicals by adding different functional groups into an aromatic ring

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lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group

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