Determination of Vitamin C Concentration by Titration

(Redox Titration Using Iodate Solution)

Introduction

This method determines the vitamin C concentration in a solution by a redox titration with potassium iodate in the presence of potassium iodide. Vitamin C, more properly called ascorbic acid, is an essential antioxidant needed by the human body (see additional notes).

When lodate lons (IO, 1) are added to an acidic solution containing lodide ions (I'), an exidation-reduction reaction occurs:

- the lodate ions are reduced to form lodine

$$10_{1}^{-} + 6 \text{ H}^{+} + 5 \text{ e}^{-} \rightarrow V_{2} I_{3} + 3 \text{ H}_{3} O$$

- while the lodide ions are oxidised to form lodine.

$$2.1^{\circ} \longrightarrow 1_{2} + 2.6^{\circ}$$

Combining these half-equations demonstrates the reaction between iodate and iodide

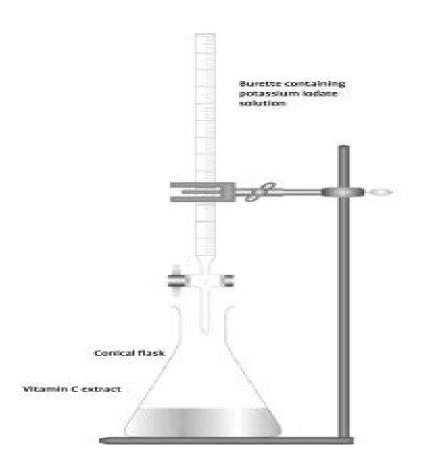
It is the iodine formed by this reaction that exidises the ascorbic acid to dehydroascorbic acid as the iodine is reduced to iodide ions.

ascorbic acid + I₂ → 2 I⁻ + dehydroascorbic acid

Due to this reaction the lodine formed is immediately reduced to lodide as long as there is any ascorbic acid present. Once all the ascorbic acid has been oxidised, the excess lodine is free to react with the starch indicator, forming the blue-black starch-lodine complex. This is the endpoint of the titration.

The method is suitable for use with Vitamin C tablets, fresh or packaged fruit juices and solid fruits and vegetables.

NB: This method is really the same as titrating ascorbic acid directly with iodine solution (see Vitamin C Method using Iodine). However, this method is more reliable as the potassium iodate solution is more stable than iodine as a primary standard.



Determination Of Vitamin C Concentration By Titration

AW Rasmussen

Determination Of Vitamin C Concentration By Titration:

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