Review for Ch. 8 Test

8.1 Describing Chemical Reactions

 know how to write <u>word equations</u> (i.e. you have be able to name compounds, including acids, ionic, and covalent compounds from Ch. 7)

- know how to write a formula equation

 What 2 things do coefficients represent in an equation? Relative number of moles of a substance; relative number of molecules needed for a reaction to occur

- know common symbols used in a reaction

- know 4 evidences of a chemical reaction:

1. energy in the form of heat of light

2. production of a gas

3. change in color

4. formation of a precipitate

 know the <u>Law of conservation of matter</u>; matter can neither be created nor destroyed in a chemical reaction (which means, mass of reactants must equal the mass of products)

- know what a chemical equation represents: relative amounts of reactants and products

- know how to balance chemical equations (combustion reactions are particularly tricky, practice!)

Complete the equation and write a word equation for the following reaction: ZnO + FeSO₄ ⇒

ZnO + FeSO₄ ===> ZnSO₄ + FeO

Zinc oxide reacts with iron (II) sulfate to yield zinc sulfate and iron (II) oxide

2. Write a formula equation for the reaction of a solution of AgCl with potassium metal:

AgCl (aq) + K (s) - KCl (aq) + Ag (s)

3. What elements are always diatomic gases at room temperature?

All of the halogens (Group 17), H2, N2, O2

4. Can you balance these reactions:

a) 2 C₂H₆ (g) + 7 O₂ (g) - 4 CO₂ (g) + 6 H₂O (g)

b) 2 Cu₂S (s) + 3 O₂ (g) = 2 Cu₂O (s) + 2 SO₂ (g)

c) 2 Al(s) + 3 Pb(NO₃)₂ (aq) = 2 Al(NO₃)₃ (aq) + 3 Pb(s)

Describing Chemical Reactions Section Review Answer Key

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