

1.36 Determine the mass of air in a  $2 \text{ m}^3$  tank if the air is at room temperature,  $20^\circ\text{C}$ , and the absolute pressure within the tank is  $200 \text{ kPa}$  (abs).

$$m = \rho V \text{ where } V = 2 \text{ m}^3 \text{ and}$$

$$\rho = p/RT \text{ with } T = 20^\circ\text{C} = (20 + 273) \text{ K} = 293 \text{ K}$$

$$\text{and } p = 200 \text{ kPa} = 200 \times 10^3 \frac{\text{N}}{\text{m}^2}$$

Thus,

$$\rho = (200 \times 10^3 \frac{\text{N}}{\text{m}^2}) / [(2.869 \times 10^2 \frac{\text{N}\cdot\text{m}}{\text{kg}\cdot\text{K}})(293 \text{ K})]$$

$$= 2.38 \frac{\text{kg}}{\text{m}^3}$$

Hence,

$$m = \rho V = 2.38 \frac{\text{kg}}{\text{m}^3} (2 \text{ m}^3) = \underline{\underline{4.76 \text{ kg}}}$$

# Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual

**Bruce R. Munson, Donald F.  
Young, Theodore H. Okiishi**



## **Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual :**

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## **Table of Contents Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual**

1. Understanding the eBook Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - The Rise of Digital Reading Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Advantages of eBooks Over Traditional Books
2. Identifying Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - User-Friendly Interface
4. Exploring eBook Recommendations from Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Personalized Recommendations
  - Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual User Reviews and Ratings
  - Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual and Bestseller Lists
5. Accessing Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual Free and Paid eBooks
  - Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual Public Domain eBooks
  - Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual eBook Subscription Services
  - Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual Budget-Friendly Options
6. Navigating Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual eBook Formats
  - ePub, PDF, MOBI, and More
  - Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual Compatibility with Devices
  - Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Highlighting and Note-Taking Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Interactive Elements Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual



8. Staying Engaged with Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
9. Balancing eBooks and Physical Books Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Setting Reading Goals Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Fact-Checking eBook Content of Fundamentals Of Fluid Mechanics 6th Edition Solutions Manual
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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