

## Chemistry – Unit 6 Worksheet 1 – Electrolysis

15. A sealed spring-piston cylinder containing hydrogen gas also contains water. The pressure of the gas in the cylinder before the temperature is  $20^{\circ}\text{C}$ . The water is left in the room, and the temperature of the gas increases to  $30^{\circ}\text{C}$ . What is the pressure in the gas?

using the combined gas law:  $P_1V_1/T_1 = P_2V_2/T_2$

using that the ratio pressure to temperature

using that the combined gas law:  $P_1V_1/T_1 = P_2V_2/T_2$

and therefore is 1.5 times

the pressure is 30

the pressure

**30 kPa**

16. A 100 mL sample of hydrogen gas with a pressure of 100 kPa is heated to  $30^{\circ}\text{C}$ . The pressure of the gas is 150 kPa. What is the volume of the gas?

the combined gas law:  $P_1V_1/T_1 = P_2V_2/T_2$

the combined gas law

the combined gas law:  $P_1V_1/T_1 = P_2V_2/T_2$

the combined gas law:  $P_1V_1/T_1 = P_2V_2/T_2$  or  $100 \times 100 = 150 \times V_2$  **66.7 mL**

17. A sealed cylinder of gas is heated to  $30^{\circ}\text{C}$ . The pressure of the gas is 100 kPa. The gas is heated to  $40^{\circ}\text{C}$ . What is the pressure of the gas?

the combined gas law

the combined gas law:  $P_1V_1/T_1 = P_2V_2/T_2$

the combined gas law

**133 kPa**

# Modeling Chemistry U9 Ws 2 V2 Answers

**Arthur James Wells**



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**The British National Bibliography** Arthur James Wells, 1979  
*Theory*, 1977

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