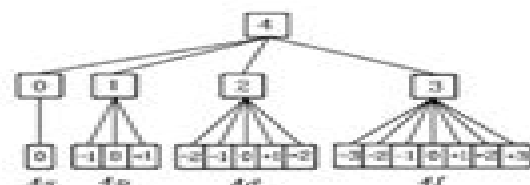


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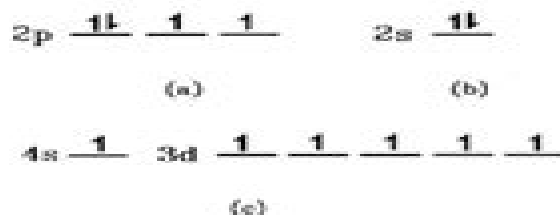
## References

- (a) Surface where electron probability = 0.
- (b) No two electrons in an atom can have the same 4 quantum numbers.
- (c) Attraction into a magnetic field by an unpaired electron.

12



- 1.5  $5p$ .
- 1.7 Size of an orbital.
- 1.9 With parallel spins there is zero probability that the electrons will occupy the same volume of space.
- 1.11 (a)  $[\text{Ne}]3s^1$ ; (b)  $[\text{Ar}]4s^2 3d^8$ ; (c)  $[\text{Ar}]4s^1 3d^{10}$ .
- 1.13 (a)  $[\text{Ar}]$ ; (b)  $[\text{Ar}]$ ; (c)  $[\text{Ar}]3d^9$ .
- 1.15  $1+$  and  $3+$ . Configuration of  $[\text{Xe}]6s^2 4f^{14} 5d^1 0 6p^1$ .
- 1.17  $1+$ . Configuration of  $[\text{Kr}]5s^1 4d^{10}$ .
- 1.19



- 1.21 E113:  $[\text{Rn}]7s^25f^{14}6d^{10}7p^1$   
 E113 as +1 ion:  $[\text{Rn}]7s^25f^{14}6d^{10}$   
 E113 as +3 ion:  $[\text{Rn}]5f^{14}6d^{10}$

## Beyond the Basics

- 1.23 9, 5, 121.
- 1.25 There are seven *f*-orbitals. There are two separate ways of depicting them and designating them: the general set and the cubic set.
- 1.27 Hydrogen heads the alkali metal group. Helium heads the alkaline earth metal group.

## Chapter 3

## Discussion

- 2.1 (a) Lanthanum through lutetium.  
(b) Apparent radius of an atom in non-bonded contact with another.  
(c) Actual nuclear charge experienced by an electron.
- 2.3 Argon did not fit into any of the then-known groups. Because the table was based on measured atomic mass, argon should have been placed between potassium and calcium.
- 2.5 The long form correctly depicts the order of elements; but the table becomes very elongated.
- 2.7 The *-ium* ending indicates a metal. The ending *-ow* has been used for non-metals.
- 2.9 With nuclei up to 26 protons, nuclear fusion is an exothermic process.
- 2.11 (a) Lead (b) technetium (c) bromine.
- 2.13 Sodium, because it has an odd number of protons.
- 2.15 50.
- 2.17 (a) Several nonmetals have metallic luster.  
(b) Diamond has the highest thermal conductivity of all substances.  
(c) Graphite is a good electrical conductor in two dimensions.
- 2.19 Potassium. As the effective atomic charge on the outermost electrons increases across a period.
- 2.21 The effective nuclear charge on the *4p* electrons will be increased.

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