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## Flame Tests

### Atomic Emission and Electron Energy Levels

#### Introduction

Just as a fingerprint is unique to each person, the color of light emitted by an element heated in a flame is also unique to each element. In this experiment, the characteristic color of light emitted by calcium, copper, lithium, potassium, sodium, and strontium ions will be observed.

#### Background

When a substance is heated in a flame, the atoms absorb energy from the flame. This absorbed energy allows the electrons to be promoted to excited energy levels. From these excited energy levels, there is a natural tendency for the electrons to make a transition or drop back down to the ground state. When an electron makes a transition from a higher energy level to a lower energy level, a particle of light called a photon is emitted (see Figure 1). Both the absorption and emission of energy are quantized—only certain energy levels are allowed.

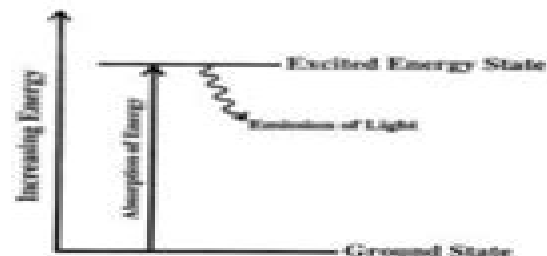


Figure 1. Absorption and Emission of Light.

An electron may drop all the way back down to the ground state in a single step, emitting a photon in the process. Alternatively, an electron may drop back down to the ground state in a series of smaller steps, emitting a photon with each step. In either case, the energy of each emitted photon is equal to the difference in energy between the excited state and the state to which the electron relaxes. The energy of the emitted photon determines the color of light observed in the flame. The flame color may be described in terms of its wavelength, and Equation 1 may be used to calculate the energy of the emitted photon.

$$\Delta E = \frac{hc}{\lambda} \quad \text{Equation 1}$$

$\Delta E$  is the difference in energy between the two energy levels in joules (J),  
 $h$  is Planck's constant ( $h = 6.626 \times 10^{-34}$  J sec)  
 $c$  is the speed of light ( $c = 3.0 \times 10^8$  m/sec)  
 $\lambda$  (lambda) is the wavelength of light in meters. The wavelengths of visible light are given in units of nanometers ( $1 \text{ m} = 1 \times 10^9 \text{ nm}$ ). See table 1 on the following page.

The color of light observed when a substance is heated in a flame varies from one substance to another. Because each element has a different spacing of electron energy levels, the possible electron transitions for a given substance are unique. Therefore, the difference in energy, between energy levels, the exact energy of the emitted photon, and the corresponding wavelength and color are unique to

# Flame Test Atomic Emission And Electron Energy Levels Answers

**Rachel S Tattersall**



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NBS Special Publication, 1968

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CRC Handbook of Basic Tables for Chemical Analysis Thomas J. Bruno, Paris D.N. Svoronos, 2020-07-30 Researchers in chemistry chemical engineering pharmaceutical science forensics and environmental science make routine use of chemical analysis but the information these researchers need is often scattered in different sources and difficult to access The CRC Handbook of Basic Tables for Chemical Analysis Data Driven Methods and Interpretation Fourth Edition is a one stop reference that presents updated data in a handy format specifically designed for use when reaching a decision point in designing an analysis or interpreting results This new edition offers expanded coverage of calibration and uncertainty and continues to include the critical information scientists rely on to perform accurate analysis Enhancements to the Fourth Edition Compiles a huge array of useful and important data into a single convenient source Explanatory text provides context for data and guidelines on applications Coalesces information from several different fields Provides information on the most useful wet chemistry methods as well as instrumental techniques with an expanded discussion of laboratory safety Contains information of historical importance necessary to interpret the literature and understand current methodology Unmatched in its coverage of the range of information scientists need in the lab this resource will be referred to again and again by practitioners who need quick easy access to the data that forms the basis for experimentation and analysis

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