

## 7.1 Fossiele brandstoffen

## A 1

- a. Aardgas bestaat voornamelijk uit methaan.  
 b. De verbrandingswarmte van aardgas is de hoeveelheid warmte die vrijkomt bij de verbranding van 1 kg of 1 m<sup>3</sup> aardgas.

A 2 ☒

- a. Gefractioneerde destillatie, die aardolie wordt gescheiden in een aantal fracties, mengsels van stoffen met een kookpunt binnen bepaalde grenzen.  
 b. Nee, de aanwezige verbindingen in aardolie worden niet ontreed. Destillatie is geen chemische reactie.  
 c. Tijdens het koken worden van grote moleculen kleinere molecuulen gemaakt. Er worden dus atoombindingen verbroken. Het koken is wel een chemische reactie.  
 d. Bij koken worden er atoombindingen verbroken om kleinere (ook waterstof)moleculen te maken uit grote koolwaterstofmoleculen. Bij reformen worden die koolwaterstofmoleculen omgezet in andere koolwaterstofmoleculen die niet per se kleiner zijn.

## A 3

Bijvoorbeeld: gevelisolatie (spouwmuurisolatie), dubbele beglazing, hoogrendementsketels voor de centrale verwarming en auto's met zuinige motoren.

## E 4

- a.  $C_{10}H_{22} \rightarrow C_7H_{16} + C_3H_8$   
 b.  $H_3C-(CH_2)_6-CH_3 \rightarrow H_3C-CH-(CH_2)_4-CH_3 + 4,5-C_4H_{10}-CH_3$   
 c. decaan  $\rightarrow$  1-buteen + hexaan  
 $C_{10}H_{22} \rightarrow C_4H_8 + C_6H_{14}$   
 $H_3C-(CH_2)_6-CH_3 \rightarrow H_3C-CH=CH-CH_2-CH_3 + H_3C-(CH_2)_4-CH_3$

## E 5

- a.  $2 C_8H_{18}(l) + 25 O_2(g) \rightarrow 16 CO_2(g) + 18 H_2O(l)$   
 b. De verbranding is exotherm, want er moet energie vrijkomen om de auto te laten rijden.  
 c. Zie figuur 7.1.

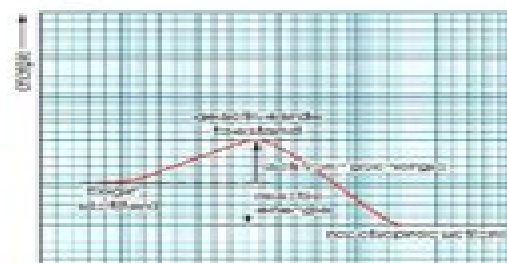


Fig. 7.1

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**M. A. W. Kabel-van den Brand**



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*Chemie overal vwo NG/NT 1* M. A. W. Kabel-van den Brand,2005

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