An automobile antifreeze mixture is made by mixing 100 mL of ethylene glycol, (CH<sub>2</sub>OH)<sub>2</sub> (p= 1.114 g/mL) and 100 mL of water (p= 1.00 g/mL) at 20 °C. The density of the mixture is 1.070 g/mL. What is the mass of solution? Leave your answer in three significant figures. \*

- O 111 g
- 56.8 g
- O 211 g
- O 100 g

# **An Antifreeze Solution Is Made By Mixing Ethylene Glycol**

**TD Snyder** 

### An Antifreeze Solution Is Made By Mixing Ethylene Glycol:

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Journal of Hydrology, New Zealand, 1979

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