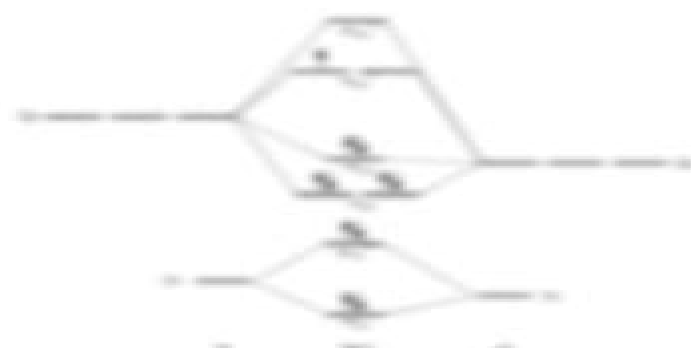


9.5	Bond Order (Figures 9.7 and 9.7c)	Bond Distance (pm)	Unpaired Electrons
$\text{O}_2$	2	120	2
$\text{O}_2^+$	2.5	121.4	1
$\text{O}_2^-$	1.5	149 (very long)	3
$\text{O}_2^{2-}$	3	126.7	0

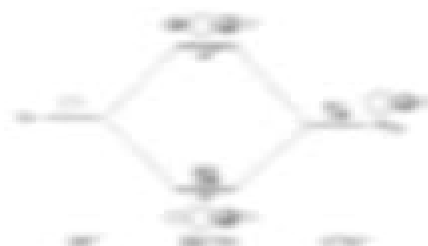
The bond distance in  $\text{O}_2^+$  is very close to the expected bond distance for a diatomic with 12 valence electrons, as shown in Figure 9.8.

- 9.6 The energy level patterns would be similar to the one shown in Figure 9.5, with the bonding orbitals for 1s and 2s orbitals (1s and 2s). All molecular orbitals except the highest would be occupied by electron pairs, and the highest orbital (i.e.,  $\pi^*$ ) would be singly occupied, giving a bond order of 1.5. Because the bond in  $\text{He}_2^+$  would be weaker than in  $\text{O}_2^+$ , the He-He distance would be expected to be longer (calculated to be  $\sim 100$  pm, see the references).

- 9.7
- The energy level diagram for  $\text{NiO}$  is on the right. The odd electrons are in a  $\pi_g^*$  orbital.
  - O is more electronegative than Ni, so its orbitals are slightly lower in energy. The bonding orbitals are slightly more concentrated on O.
  - The bond order is 2.5, with one unpaired electron.
  - $\text{Ni}^{2+}$  Bond order = 2  
 shortest bond (190 pm)  
 $\text{NiO}$  Bond order = 2.5  
 intermediate (207 pm)  
 $\text{Ni}^{+}$  Bond order = 2  
 longest bond (227 pm), two electrons in antibonding orbitals.



- 9.8
- The  $\text{CN}^-$  energy level diagram is similar to that of  $\text{N}_2$  (Problem 9.7) without the antibonding  $\sigma^*$  electrons.
  - The bond order is three, with no unpaired electrons.
  - The HOMO is the  $\pi_g$  orbital, which can interact with the 1s of the H<sup>+</sup> as in the diagram on right. The bonding orbital has an energy lower than of the  $\pi$  orbitals; the antibonding orbital has more the highest energy orbital.



# Answers To Inorganic Chemistry Miessler

**Anthony S. Fauci**



## **Answers To Inorganic Chemistry Miessler:**

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In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Answers To Inorganic Chemistry Miessler** a literary masterpiece penned by way of a renowned author, readers embark on a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of people who partake in its reading experience.

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### **Table of Contents Answers To Inorganic Chemistry Miessler**

1. Understanding the eBook Answers To Inorganic Chemistry Miessler
  - The Rise of Digital Reading Answers To Inorganic Chemistry Miessler
  - Advantages of eBooks Over Traditional Books
2. Identifying Answers To Inorganic Chemistry Miessler
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers To Inorganic Chemistry Miessler
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answers To Inorganic Chemistry Miessler
  - Personalized Recommendations
  - Answers To Inorganic Chemistry Miessler User Reviews and Ratings
  - Answers To Inorganic Chemistry Miessler and Bestseller Lists

5. Accessing Answers To Inorganic Chemistry Miessler Free and Paid eBooks
  - Answers To Inorganic Chemistry Miessler Public Domain eBooks
  - Answers To Inorganic Chemistry Miessler eBook Subscription Services
  - Answers To Inorganic Chemistry Miessler Budget-Friendly Options
6. Navigating Answers To Inorganic Chemistry Miessler eBook Formats
  - ePub, PDF, MOBI, and More
  - Answers To Inorganic Chemistry Miessler Compatibility with Devices
  - Answers To Inorganic Chemistry Miessler Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answers To Inorganic Chemistry Miessler
  - Highlighting and Note-Taking Answers To Inorganic Chemistry Miessler
  - Interactive Elements Answers To Inorganic Chemistry Miessler
8. Staying Engaged with Answers To Inorganic Chemistry Miessler
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answers To Inorganic Chemistry Miessler
9. Balancing eBooks and Physical Books Answers To Inorganic Chemistry Miessler
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answers To Inorganic Chemistry Miessler
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answers To Inorganic Chemistry Miessler
  - Setting Reading Goals Answers To Inorganic Chemistry Miessler
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers To Inorganic Chemistry Miessler
  - Fact-Checking eBook Content of Answers To Inorganic Chemistry Miessler
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
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