# Answers

# Chapter 1

#### Exercises

- (a) CuCO<sub>2</sub> → CuO + CO<sub>2</sub>
  - (b) 2Mg + O<sub>5</sub> → 2MgO
  - (c) H,SO, + 2NaOH -+ Na,SO, + 2H,O
  - (d) N<sub>s</sub> + 3H<sub>s</sub> → 2NH<sub>s</sub>
  - (e) CH<sub>4</sub> + 2O<sub>2</sub> → CO<sub>2</sub> + 2H<sub>2</sub>O
- 2 (a) 2K + 2H,O → 2KOH + H,
  - (b) C,H,OH + 3O, → 2CO, + 3H,O
  - (c) Cl, + 2KI → 2KCI + 1,
  - (d) 4CrO, → 2Cr,O, + 3O,
  - (e) Fe<sub>3</sub>O<sub>5</sub> + 3C → 3CO + 2Fe
- 3 (a) 2C<sub>4</sub>H<sub>10</sub> + 13O<sub>2</sub> → 8CO<sub>2</sub> + 10H<sub>2</sub>O
  - (b) 4NH<sub>2</sub> + 5O<sub>2</sub> → 4NO + 6H<sub>2</sub>O
  - (c) 3Cu + 8HNO<sub>5</sub> → 3Cu(NO<sub>5</sub>)<sub>2</sub> + 2NO + 4H<sub>2</sub>O
  - (d) 6H,O, + 2N,H, → 2N, + 10H,O + O,
  - (e) 4C<sub>2</sub>H<sub>2</sub>N + 15O<sub>2</sub> → 8CO<sub>2</sub> + 14H<sub>2</sub>O + 2N<sub>2</sub>
- 4 (a) Sand and water: heterogeneous
  - (b) Smoke: heterogeneous
  - (c) Sugar and water: homogeneous
  - (d) Salt and iron filings; heterogeneous
  - (e) Ethanol and water: homogeneous
  - (f) Steel: homogeneous
- 5 (a) 2KNO<sub>3</sub>(s) → 2KNO<sub>3</sub>(s) + O<sub>3</sub>(g)
  - (b) CaCO₂(s) + H₂SO₂(aq) → CaSO₂(s) + CO₂(g) + H₂O(l)
  - (c)  $2\text{Li}(s) + 2\text{H}_{y}\text{O}(t) \rightarrow 2\text{LiOH}(aq) + \text{H}_{y}(q)$
  - (d) Pb(NO<sub>J</sub>(aq) + 2NaCl(aq) → PbCl<sub>J</sub>(s) + 2NaNO<sub>J</sub>(aq)
  - (e) 2C<sub>2</sub>H<sub>2</sub>(g) + 9O<sub>2</sub>(g) → 6CO<sub>2</sub>(g) + 6H<sub>2</sub>O(l)

In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid, as shown here.

- 6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)
- 7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid, and therefore that ice would sink in water.
- 8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.
- 9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase.
- Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E-D on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.
- 11 8
- 12 AW 01\_EX\_12

# **Answers For Ib Chemistry Pearson Truklyore**

**Wolfgang Guggemos** 

**Answers For Ib Chemistry Pearson Truklyore:** 

#### Answers For Ib Chemistry Pearson Truklyore Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the power of words has be much more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Answers For Ib Chemistry Pearson Truklyore**, a literary masterpiece that delves deep in to the significance of words and their impact on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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