

Answers to The Oceans End of Unit Test

Q	Answer with marks	Marking suggestions
1(a)	CO ₂ is a gas (1); which has more entropy as there are more ways of arranging its molecules (1)	
1(b)	$\Delta S^{\circ}_{\text{sys}} = 214 + 38 - 93$ (1); $= +159 \text{ J K}^{-1} \text{ mol}^{-1}$ (1), including sign and units	
1(c) (i)	$-180 \times 1000/300$ (1); $= -600 \text{ J K}^{-1} \text{ mol}^{-1}$ (1), including sign and units	
1(c) (ii)	-150 (1) ($\text{J K}^{-1} \text{ mol}^{-1}$)	
1(d)	$\Delta S^{\circ}_{\text{tot}} = \Delta S^{\circ}_{\text{sys}} + \Delta S^{\circ}_{\text{sur}}$ This is negative at 300 K and positive at 1200 K (1); Positive values means reactions are possible (1)	
1(e) (i)	 Top line correct, labelled ΔH_f or enthalpy change of formation (1); Rest of cycle correct, arrows labelled ΔH or some attempt to assign specific enthalpy changes (1)	Or alternative cycle drawn as an enthalpy level diagram
1(e) (ii)	Lattice enthalpy $= \Delta H_3 = \Delta H_1 - \Delta H_2 - \Delta H_4$ (1) $= -635 - (178 + 249) - (596 + 1152 - 147 + 753)$ $= -3416 \text{ kJ mol}^{-1}$ (1)	Units not essential for mark
1(f) (i)	At the bottom of the lake the pressure is higher (1); this means the equilibrium will move to the left, increasing the concentration of CO ₂ (aq) (1)	
1(f) (ii)	Deep oceans can absorb a lot of CO ₂ (1); this acts as a 'sink' for the extra CO ₂ we are producing (1)	

Q	Answer with marks	Marking suggestions
2(a)	Ammonia has only (an average) of one hydrogen bond per molecule (1); water has two (1); greater intermolecular forces mean higher boiling point (1)	
2(b)	Ammonia would be less efficient (1); as it has a lower specific heating capacity than water (1); thus less energy would be transferred per mole/gram of ammonia (1)	

Q	Answer with marks	Marking suggestions
3(a) (i)	It is an equilibrium (1); so the benzoic acid is not fully reacted/ ionised (1)	
3(a) (ii)	$K_a = \frac{[\text{C}_6\text{H}_5\text{COO}^-][\text{H}_3\text{O}^+]}{[\text{C}_6\text{H}_5\text{COOH}]}$ (1) for products on nominator; reactants on denominator; (1) completely correct	
3(b) (i)	$-\lg[\text{H}^+]/-\lg[\text{H}_3\text{O}^+]$ (1)	
3(b) (ii)	$[\text{H}_3\text{O}^+]^2 = (1 \times 10^{-4}) \times 0.010 = 1 \times 10^{-6}$ (1) $[\text{H}_3\text{O}^+] = 1 \times 10^{-3} \text{ mol dm}^{-3}$ (1) $\text{pH} = 3$ (1)	
3(b) (iii)	$\text{pH} = 2$ (1)	
3(b) (iv)	HCl is fully ionised (1); (benzoic acid is partially ionised) Hence the concentration of $[\text{H}_3\text{O}^+]$ is lower (1)	

Answers To The Atmosphere End Of Unit Test Benjamin Mills

Terry C. Jones



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In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Answers To The Atmosphere End Of Unit Test Benjamin Mills**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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