Experiment 25

Observations and Calculations: pH: Buffers and Their Properties

A. Determination of pH by the Use of Acid-Base Indicators

	Color with 0.1 M Solution of		
Indicator IICI	NaH PO.	THUMO	ZaSO.
DIVE	Doru word t	The second secon	DUPPLE
RED			MENON
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			eld-prange
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	DIVE	PLE DONE WITH L PLE YELLOW PLE YELLOW PLE PROUNT YELLOW YELLOW ACCELO	BILL DOOK PUTTLE DOOK PUTTLE RED YELLOW STATE PLE YELLOW PLE BROWN PORF PUTTLE YELLOW ACCELO YELLOW

Circle the observation(s) for each solution that was (were) most useful in estimating the pH range.

B. Measurement of the pH of Some Typical Solutions

Record the pH and the color observed with bromcresol green for each of the 0.1 M solutions that were tested.

pH A.9A 10.97 7.17 NaHSO. 10.97 T.17 LIFT

For any two solutions having a pH less than 6 or greater than 8, write a net ionic equation to explain qualitatively why the solution has that pH.

Solution NA 2 CO3 Equation $CO3^2 + H20 \Rightarrow HCO3 + OH^2$ Solution NA 11504 Equation $H509^- + H20 \Rightarrow (504)^2 + H30^+$

Explain why the color observed with bromcresol green for each of the four solutions is reasonable, given the pH.

C. Some Properties of Buffers

Buffer system selected HCO3 - CO3 HB is hydrogen cortonate - (or beneft (name the acid)

1. pH of buffer 9-29 [H*] 5.24 × 10 M pK, (by Eq. 4b)

2. pH of diluted buffer 9.24 [H*] 5.75 × 10 M pK.

pH after addition of 5 drops NaOH

pH after addition of 5 drops HCl

Comment on your observations in Parts 1 and 2.

Ph Properties Of Buffer Solutions Pre Lab Answers

Daniela Niemeyer

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proceedings primarily abstracts of papers presented at the 10th 53rd annual meetings and the 1948 56 fall meetings $\underline{\text{Excerpta Medica}}$, 1963

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