

**Physical Chemistry Test #2**

(MM:29 | 60 min)

**Question 1 : Answer the following questions :**

[4]

- (i) The pressure is directly proportional to mole-fraction of gas in solution. This law is called \_\_\_\_\_.
- (ii) The metal which is highly reactive will have  
(a) high value of standard reduction potential.  
(b) low value of standard reduction potential.  
(c) high value of standard oxidation potential.  
(d) low value of standard oxidation potential.
- (iii) Calculate the overall order of reaction whose rate law expression is given  
 $r = k[\text{NO}]^{3/2} [\text{O}_2]^{1/2}$   
(a) 2<sup>nd</sup>                      (b) 4<sup>th</sup>                      (c) 3<sup>rd</sup>                      (d) 1<sup>st</sup>.
- (iv) **Assertion :** Vapour pressure of a solution increases with increase in temperature.  
**Reason :** Rate of evaporation exceeds rate of condensation.  
(a) Both assertion and reason are true and reason is the correct explanation of assertion.  
(b) Both assertion and reason are true but reason is not the correct explanation for assertion.  
(c) Assertion is true but reason is false.  
(d) Assertion is false but reason is true.

**Question 2**

[2]

A solution is prepared by dissolving 9.25g of non-volatile solute in 450ml of water. It has an osmotic pressure of 350mm of Hg at 27°C. Assuming the solute is non-electrolyte, determine its molecular mass. ( $R = 0.0821 \text{ lit atm K}^{-1} \text{ mol}^{-1}$ )

**Question 3**

[2]

Calculate the mass of silver deposited at cathode when a current of 2 amperes is passed through a solution of  $\text{AgNO}_3$  for 15 minutes. (at. wt. of Ag = 108, 1 F = 96,500 C)

**Question 4**

[2]

When 0.4g of acetic acid is dissolved in 40g of benzene, the freezing point of the solution is lowered by 0.45K. Calculate the degree of association of acetic acid. Acetic acid forms dimer when dissolved in benzene. ( $K_f$  for benzene =  $5.12 \text{ K kg mol}^{-1}$ , at. wt. C = 12, H = 1, O = 16)

**Question 5**

[3]

25% of a first order reaction is completed in 30 minutes. Calculate the time taken in minutes for the reaction to go to 90% completion.

**Question 6**

[3]

Define the following terms:

(1) Raoult's law for elevation in boiling point.

(2) Kohlrausch's Law

# Physical Chemistry Exam Quiz Homework With Solution

**JA Banks**



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