

CLASS 10TH SCIENCE



PERIODIC CLASSIFICATION OF ELEMENTS

A standard periodic table of elements, color-coded by groups. The elements are arranged in rows and columns, with their chemical symbols and atomic numbers. The table includes elements from Hydrogen (H) to Oganesson (Og).

1	2		3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18																				
H	He		Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar																				
			Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr										
			K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
			Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn	
			Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn	Fr	Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr		
			Fr	Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr																			

Legend:

- 1. Hydrogen (H)
- 2. Helium (He)
- 3. Lithium (Li)
- 4. Beryllium (Be)
- 5. Boron (B)
- 6. Carbon (C)
- 7. Nitrogen (N)
- 8. Oxygen (O)
- 9. Fluorine (F)
- 10. Neon (Ne)
- 11. Sodium (Na)
- 12. Magnesium (Mg)
- 13. Aluminum (Al)
- 14. Silicon (Si)
- 15. Phosphorus (P)
- 16. Sulfur (S)
- 17. Chlorine (Cl)
- 18. Argon (Ar)
- 19. Potassium (K)
- 20. Calcium (Ca)
- 21. Scandium (Sc)
- 22. Titanium (Ti)
- 23. Vanadium (V)
- 24. Chromium (Cr)
- 25. Manganese (Mn)
- 26. Iron (Fe)
- 27. Cobalt (Co)
- 28. Nickel (Ni)
- 29. Copper (Cu)
- 30. Zinc (Zn)
- 31. Gallium (Ga)
- 32. Germanium (Ge)
- 33. Arsenic (As)
- 34. Selenium (Se)
- 35. Bromine (Br)
- 36. Krypton (Kr)
- 37. Rubidium (Rb)
- 38. Strontium (Sr)
- 39. Yttrium (Y)
- 40. Zirconium (Zr)
- 41. Niobium (Nb)
- 42. Molybdenum (Mo)
- 43. Technetium (Tc)
- 44. Ruthenium (Ru)
- 45. Rhodium (Rh)
- 46. Palladium (Pd)
- 47. Silver (Ag)
- 48. Cadmium (Cd)
- 49. Indium (In)
- 50. Tin (Sn)
- 51. Antimony (Sb)
- 52. Tellurium (Te)
- 53. Iodine (I)
- 54. Xenon (Xe)
- 55. Barium (Ba)
- 56. Lanthanum (La)
- 57. Hafnium (Hf)
- 58. Tantalum (Ta)
- 59. Tungsten (W)
- 60. Rhenium (Re)
- 61. Osmium (Os)
- 62. Iridium (Ir)
- 63. Platinum (Pt)
- 64. Gold (Au)
- 65. Mercury (Hg)
- 66. Thallium (Tl)
- 67. Lead (Pb)
- 68. Bismuth (Bi)
- 69. Polonium (Po)
- 70. Astatine (At)
- 71. Radon (Rn)
- 72. Francium (Fr)
- 73. Radium (Ra)
- 74. Actinium (Ac)
- 75. Thorium (Th)
- 76. Protactinium (Pa)
- 77. Uranium (U)
- 78. Neptunium (Np)
- 79. Plutonium (Pu)
- 80. Americium (Am)
- 81. Curium (Cm)
- 82. Berkelium (Bk)
- 83. Californium (Cf)
- 84. Einsteinium (Es)
- 85. Fermium (Fm)
- 86. Mendelevium (Md)
- 87. Nobelium (No)
- 88. Lawrencium (Lr)



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As 2019 has been declared the International Year of the Periodic Table it is appropriate that Structure and Bonding marks this anniversary with two special volumes In 1869 Dmitri Ivanovitch Mendeleev first proposed his periodic table of the elements He is given the major credit for proposing the conceptual framework used by chemists to systematically inter relate the chemical properties of the elements However the concept of periodicity evolved in distinct stages and was the culmination of work by other chemists over several decades For example Newland s Law of Octaves marked an important step in the evolution of the periodic system since it represented the first clear statement that the properties of the elements repeated after intervals of 8 Mendeleev s predictions demonstrated in an impressive manner how the periodic table could be used to predict the occurrence and properties of new elements Not all of his many predictions proved to be valid but the discovery of scandium gallium and germanium represented sufficient vindication of its utility and they cemented its enduring influence Mendeleev s periodic table was based on the atomic weights of the elements and it was another 50 years before Moseley established that it was the atomic number of the elements that was the fundamental parameter and this led to the prediction of further elements Some have suggested that the periodic table is one of the most fruitful ideas in modern science and that it is comparable to Darwin s theory of evolution by natural selection proposed at approximately the same time There is no doubt that the periodic table occupies a central position in chemistry In its modern form it is reproduced in most undergraduate inorganic textbooks and is present in almost every chemistry lecture room and classroom This second volume provides chemists with an overview of the important role played by the Periodic Table in advancing our knowledge of solid state and bioinorganic chemistry It also illustrates how it has been used to fine tune the properties of compounds which have found commercial applications in catalysis electronics ceramics and in medicinal chemistry

The Periodic Table of Elements Compendium Helen Eccles,1994 This book provides descriptions of 103 elements their properties occurrences and uses plus a summary of periodic table patterns **Periodic Table of Elements** Macmillan Education, Limited,1997-11-03 A wallchart of the periodic table of elements The Elements Adrian Dingle,2017

Periodic Table of the Elements Ekkehard Fluck,Klaus G. Heumann,1986-07-09

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