

[4101] - 101

F.Y. B.A.

COMPULSORY ENGLISH

Text Prescribed : REFLECTIONS - I

Ed. by P. Bhaskaran Nair, Geetha Rajeevan & G. Radhakrishna Pillai  
(1014) (2008 Pattern)

Time : 3 Hours]

[Max. Marks : 80

Instructions to the candidates:-

- 1) All questions are compulsory.
- 2) Figures to the right indicate full marks.

**Q1)** Attempt any two of the following : [16]

- a) What message does Rabindranath Tagore give in his poem 'Leave this Chanting and Singing and Telling of Beads'?
- b) Comment on the changing decisions of Fedor Sigaev in 'The Avenger'.
- c) Why did Wasserkopf want a refund of the tuition fee he had paid to the school?
- d) How does Sylvia plath describe truthfulness of the mirror?

**Q2)** Attempt any four of the following: [16]

- a) What were the different ways in which Nicola and Jacopo earned money?
- b) How many parts does a human being act in a single play called life? Discuss with reference to 'All the World's a Stage'.
- c) How did the Conjuror damage Quick Man's things?
- d) How does uncle Podger engage his family members and servants to hang a picture?
- e) Write a note on the pathetic condition of mother in 'The Affliction of Margaret'.
- f) What is the central idea of the poem 'Gather ye Rosebuds while ye May'?
- g) Write a brief note on the effects caused by tsunami as depicted by Amitav Ghosh in 'The Town by the Sea'.

# Ib English B Hl Past Papers 2012

**Léo Pomar**



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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic

eg 1  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$  for which  $\Delta H = -57 \text{ kJ mol}^{-1}$  eg 2  $\text{HNO}_3\text{(aq)} + \text{KOH(aq)} \rightarrow \text{KNO}_3\text{(aq)} + \text{H}_2\text{O(l)}$  for which  $\Delta H = -57 \text{ kJ mol}^{-1}$   
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