



## Topic 19.1: Organic chemistry – Fundamentals of organic chemistry

Organic chemistry focuses on the chemistry of compounds containing carbon.

• **Understanding:** A homologous series is a series of compounds of the same family, with the same general formula, which differ from each other by a common structural unit.

- **Homologous series:** series of compound of the same family, with the same general formula, which differ by a common structural unit
  - Common structure unit: often  $\text{CH}_2$  (methylene) group
  - Same family: has the same functional group
  - Properties: homologous series has similar physical and chemical properties

• **Understanding:** Structural formulas can be represented in full and condensed format.

	Full structural formula	Condensed structural formula	Skeletal structural formula
Example		$\text{CH}_3\text{—CH}_2\text{—CH}_2\text{—CH}_3$  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	
Description	Two-dimensional representation showing all atoms, their relative positions and bonds	Two-dimensional representation showing all atoms and their relative position	Two dimensional representation of a compound where each vertex and end of line represent a carbon atom, and functional group is added

• **Understanding:** Structural isomers are compounds with the same molecular formula but different arrangements of atoms.

- **Structural isomers:** compounds with same molecular formula with different arrangement of atoms
  - Properties: structural isomers have different chemical and physical properties (melting point, boiling point etc.)

• **Understanding:** Functional groups are the reactive parts of molecules.

- **Functional groups:** reactive parts of a molecule that is responsible for the molecule's chemical property
  - Naming: the position of the function group is identified by the number of carbon to which it is attached

• **Understanding:** Saturated compounds contain single bonds only and unsaturated compounds contain double or triple bonds.

- **Saturated compounds:** organic compounds in which all the carbon-carbon bonds are single (e.g. **alkanes**)
- **Unsaturated compounds:** organic compound in which there is at least one carbon-carbon bond that is double or triple (e.g. **alkenes**)

• **Understanding:** Benzene is an aromatic, unsaturated hydrocarbon.

- **Aliphatic:** characterized by a straight, branched or cyclic structure
- **Aromatic:** characterized by presence of benzene ring
- **Benzene:** aromatic, unsaturated hydrocarbon in which six carbons bonds into a ring
  - **Sigma ( $\sigma$ ) bond in benzene:** benzene contains six  $\text{sp}^2$  hybridized carbon atoms bond to one other by sigma bonds
  - **Pi ( $\pi$ ) bond in benzene:** p-orbitals in carbon overlap to form a continuous  $\pi$  bond above and below the plane of carbons

• **Applications and skills:** Explanation of the trends in boiling points of members of a homologous series.

Name	Formula	Structural formula (condensed)	Boiling point ( $^{\circ}\text{C}$ )
Methane	$\text{CH}_4$	$\text{CH}_4$	-161
Ethane	$\text{C}_2\text{H}_6$	$\text{CH}_3\text{CH}_3$	-89
Propane	$\text{C}_3\text{H}_8$	$\text{CH}_3\text{CH}_2\text{CH}_3$	-42
Butane	$\text{C}_4\text{H}_{10}$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	-0.5
Pentane	$\text{C}_5\text{H}_{12}$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	36
Hexane	$\text{C}_6\text{H}_{14}$	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	69

- **Trend in boiling points:** boiling points increases as the carbon chain increases by a common structural unit
  - Trends arise from increasing intermolecular (dispersion forces) with increased chain lengths

• **Applications and skills:** Distinction between empirical, molecular and structural formulas.

- **Empirical formula:** simplest ratio of atoms (of an element) in a molecule
- **Molecular formula:** actual number of atoms (of each element) in a molecule
- **Structural formula:** representation of the type of atoms and their position relative to one other

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## **Table of Contents Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes**

1. Understanding the eBook Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - The Rise of Digital Reading Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - Advantages of eBooks Over Traditional Books
2. Identifying Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes

- Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
- Popular eBook Platforms
  - Features to Look for in an Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
- Personalized Recommendations
  - Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes User Reviews and Ratings
  - Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes and Bestseller Lists
5. Accessing Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes Free and Paid eBooks
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  - Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes eBook Subscription Services
  - Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes Budget-Friendly Options
6. Navigating Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes eBook Formats
- ePub, PDF, MOBI, and More
  - Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes Compatibility with Devices
  - Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes Enhanced eBook Features

7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - Highlighting and Note-Taking Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - Interactive Elements Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
8. Staying Engaged with Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
9. Balancing eBooks and Physical Books Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - Setting Reading Goals Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher Level Ib Chemistry Revision Notes
  - Fact-Checking eBook Content of Ib Chemistry 1 Quantitative Chemistry Revision Notes Standard And Higher

Level Ib Chemistry Revision Notes

- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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