

Quesiti e problemi

1 I nomi delle sostanze

1 Chi è stato il primo scienziato a tentare di razionalizzare i nomi delle sostanze chimiche?

Lavoisier

2 Quali nomenclature sono attualmente utilizzate?

nomenclatura tradizionale; nomenclatura di Stock;
nomenclatura IUPAC

3 Che cos'è la IUPAC e di che cosa si occupa?

Associazione internazionale con lo scopo
di razionalizzare la nomenclatura dei composti chimici.

2 Valenza e numero di ossidazione

4 Nel composto N_2O_5 , gli atomi di azoto hanno valenza 5 e quelli di ossigeno hanno valenza 2.

► Qual è il numero di ossidazione dell'azoto?

+5

► Qual è il numero di ossidazione dell'ossigeno?

-2

5 Qual è il numero di ossidazione del silicio nel composto SiH_4 ?

+4

6 Utilizzando la tabella 12.1, attribuisce il numero di ossidazione agli elementi presenti nelle seguenti formule costituite da due o tre elementi.

a) $Ca(OH)_2$ $Ca: +2; O: -2; H: +1$ b) CaO $Ca: +2; O: -2$ c) $MgSO_4$ $Mg: +2; S: +6; O: -2$ d) Fe_2O_3 $Fe: +3; O: -2$ e) H_2O_2 $H: +1; O: -1$ f) V_2O_5 $V: +5; O: -2$ g) NH_3 $N: -3; H: +1$ h) MnO_2 $Mn: +4; O: -2$ i) HCO_3^- $H: +1; C: +4; O: -2$ l) PO_4^{3-} $P: +5; O: -2$ m) $KMnO_4$ $K: +1; Mn: +7; O: -2$ n) PbI_2 $Pb: +2; I: -1$ o) KNO_3 $K: +1; N: +5; O: -2$ p) Cr_2O_3 $Cr: +3; O: -2$ q) H_4SiO_4 $H: +1; Si: +4; O: -2$ r) $SrSO_4$ $Sr: +2; S: +6; O: -2$ s) P_2O_5 $P: +5; O: -2$ t) MnO_4^{2-} $Mn: +6; O: -2$ u) Sb_2S_3 $Sb: +3; S: -2$ v) IO_3^- $I: +5; O: -2$

7 Scrivi la formula di un composto dell'azoto con numero di ossidazione -3.

 NH_3

8 Il manganese può avere numeri di ossidazione variabili tra +2 e +7. Qual è il suo n.o. nello ione manganato MnO_4^{2-} ?

+4

9 Utilizzando la tabella 12.1, attribuisce il numero di ossidazione agli elementi presenti nelle seguenti formule, costituite da uno o due elementi.

a) S_8 $S: 0$ b) H_2O $H: +1; O: -2$ c) Hg $Hg: 0$ d) SO_3 $S: +6; O: -2$ e) PH_3 $P: -3; H: +1$ f) Cl_2O_7 $Cl: +7; O: -2$ g) CH_4 $C: -4; H: +1$ h) N_2 $N: 0$ i) LiH $Li: +1; H: -1$ l) CO $C: +2; O: -2$ m) Xe $Xe: 0$ n) Al_2O_3 $Al: +3; O: -2$ o) CuO $Cu: +2; O: -2$ p) Cu_2O $Cu: +1; O: -2$ q) $AgCl$ $Ag: +1; Cl: -1$ r) Na_2S $Na: +1; S: -2$ s) F^- $F: -1$ t) SO_4^{2-} $S: +6; O: -2$ u) Ca^{2+} $Ca: +2$ v) MnO_4^- $Mn: +7; O: -2$

10 Trova i numeri di ossidazione degli elementi sottolineati.

a) Cl ClO $+1$ b) Sg H_2 -2 c) Si H_4 $+4$ d) Ag $_2O$ $+1$ e) C O_2 $+4$ f) Si Cl_4 $+4$ g) K $_2SO_4$ $+1$ h) Fe O $+2$ i) O $_3$ 0 l) H $_2SO_3$ $+4$ m) H NO_2 $+3$ n) Ba SO_4 $+2$ o) H $_2S$ -2 p) N H_3 -3

11 Quali elementi formano ioni positivi (cationi) e quali formano ioni negativi (anioni)?

In generale: i metalli tendono a trasformarsi in cationi
e i non metalli in anioni.

12 Trova i numeri di ossidazione degli elementi sottolineati.

a) Hg I_2^{2+} $+1$ b) Si F_6^{2-} $+4$ c) S $_2O_8^{2-}$ $+7$ d) Sg Cl_5^- $+4$ e) Ga Cl_3^- $+3$ f) NH $_4^+$ -3 g) N O_3^- $+5$ h) Cr $_2O_7^{2-}$ $+6$ i) Cl O_4^- $+7$ l) NH $_2^-$ -3 m) N O_2^- $+5$ n) Mg O_4^{2-} $+6$

3 Leggere e scrivere le formule più semplici

13 Scrivi la formula bruta dei composti binari che si formano dai seguenti elementi.

a) K e I KI b) S (con n.o. -2) e Cl SO_2 c) Sr e H SrH_2 d) H e S H_2S

Zanichelli Soluzioni Chimica

Anna Faresin



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