

Understanding
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Fifth Edition

Virginia Klenk

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 partly because so many compounds can act as acids or bases

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 a 25 0 ml sample of 0 100 m hydrochloric acid with 0 100 m sodium hydroxide the values of the ph measured after successive
 additions of small amounts of naoh are listed in the first column of this table and are graphed in

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 until the solution is exactly neutralised i e there is only salt and water it is possible to calculate

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 titration of an acid with a base or vice versa the indicator is a substance that can exist in two forms an

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 example below demonstrates the technique to solve a titration problem for a titration of sulfuric acid with sodium hydroxide

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web as seen in the chapter on the stoichiometry of chemical reactions titrations can be used to quantitatively analyze solutions for their acid or base concentrations in this section we will explore the underlying chemical equilibria that make acid base titrimetry a

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