

## Lab: Baking Soda Stoichiometry

**Introduction:** When baking soda ( $\text{NaHCO}_3$ ) is an ingredient in your recipe, its purpose is to make the batter rise and produce a product with a light and fluffy texture. The decomposition reaction is shown below:



Our goal will be to predict how much sodium carbonate ( $\text{Na}_2\text{CO}_3$ ) is produced when baking soda decomposes.

**Materials:** [Lab Video \(use to fill in data table\)](#)

### **Procedure:**

1. Use the lab video to complete the data table below.
2. Calculate the mass of the  $\text{NaHCO}_3$  used in the reaction and the experimental mass of  $\text{Na}_2\text{CO}_3$  that was produced. **You need to show work in the "calculation" table for how you got your answer. Typing it out is okay for the calculation table section.**
3. Answer the 2 conclusion questions at the end.

### **Data table**

Mass of crucible (g)	12.74 g
Mass of crucible and $\text{NaHCO}_3$ (g)	14.33 g
Mass of crucible and $\text{Na}_2\text{CO}_3$ (g)	13.75 g

### **Calculations Table (show work in each box)**

Mass of $\text{NaHCO}_3$ (g)	1.59 g
Experimental (actual) mass of $\text{Na}_2\text{CO}_3$ (g)	1.01 g

**Conclusion Questions:** Show all work, handwritten, using the correct steps/methods shown during class. You will need to upload an image of your work.

1. Use your starting mass of  $\text{NaHCO}_3$  and the balanced chemical equation above to calculate the mass of  $\text{Na}_2\text{CO}_3$  that **should** be produced by the decomposition reaction. This is the theoretical yield.

# Baking Soda Stoichiometry Lab Answers

**Robert Bruce Thompson**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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