

D.	2	$\text{KClO}_3$	2	$\text{KCl} +$	3	$\text{O}_2$
E.	2	$\text{Al} +$	3	$\text{CuO}$	1	$\text{Al}_2\text{O}_3 +$
F.	1	$\text{I}_2 +$	2	$\text{Na}_2\text{S}_2\text{O}_3$	2	$\text{NaI} +$
G.	6	$\text{Mg} +$	1	$\text{P}_2$	2	$\text{Mg}_3\text{P}_2$

**Activity B:**  
Classifying reactions

Get the Gizmo ready:

- Turn off Show summary and Show histograms.



**Introduction:** Chemical equations show how **compounds** and **elements** react with one another. An element is a substance consisting of one kind of atom, such as aluminum (Al) or oxygen gas ( $\text{O}_2$ ). A compound is a substance made of more than one kind of atom, such as water ( $\text{H}_2\text{O}$ ) or table salt (NaCl).

**Question: How are chemical reactions classified?**

1. **Match:** Most chemical reactions can be classified as one of four types. Using the chemical equations in the Gizmo as a guide, match the following definitions to the type of reaction.

- |   |   |
|---|---|
| B | One reactant is broken down into two or more products.  |
| E | A fuel combines with oxygen to produce carbon dioxide ( $\text{CO}_2$ ) and water ( $\text{H}_2\text{O}$ ). |
| A | Two or more reactants combine to form one product.  |
| D | Two compounds react to form two different compounds.  |
| C | A compound reacts with an element to form a new compound and a different element.                           |

- |    |                           |
|----|---------------------------|
| A. | <b>Synthesis</b>          |
| B. | <b>Decomposition</b>      |
| C. | <b>Single replacement</b> |
| D. | <b>Double replacement</b> |
| E. | <b>Combustion</b>         |

2. **Practice:** Balance each of the chemical equations below. (Some equations may already be in balance.) In the space to the right, classify the reaction as a **synthesis**, **decomposition**, **single replacement**, **double replacement**, or **combustion**.

A.	1	$\text{AgNO}_3 +$	1	$\text{KCl}$	<input type="checkbox"/>	1	$\text{AgCl} +$	1	$\text{KNO}_3$	Two replacements
B.	1	$\text{H}_2\text{O} +$	1	$\text{SO}_3$	<input type="checkbox"/>	1	$\text{H}_2\text{SO}_4$			Synthesis
C.	2	$\text{KI} +$	1	$\text{Cl}_2$	<input type="checkbox"/>	2	$\text{KCl} +$	1	$\text{I}_2$	Single replacement
D.	2	$\text{NaHCO}_3$	<input type="checkbox"/>	1	$\text{Na}_2\text{CO}_3 +$	1	$\text{H}_2\text{O} +$	1	$\text{CO}_2$	Decomposition
E.	1	$\text{Zn} +$	2	$\text{HCl}$	<input type="checkbox"/>	1	$\text{ZnCl}_2 +$	1	$\text{H}_2$	Single replacement

# Balancing Chemical Equations Gizmo Answer Key

**Jake Mariska**

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