

Lab: Baking Soda Stoichiometry

Introduction: When baking soda (NaHCO_3) is an ingredient in your recipe, its purpose is to make the batter rise and produce a product with a light and fluffy texture. The decomposition reaction is shown below:



Our goal will be to predict how much sodium carbonate (Na_2CO_3) is produced when baking soda decomposes.

Materials: [Lab Video \(use to fill in data table\)](#)

Procedure:

1. Use the lab video to complete the data table below.
2. Calculate the mass of the NaHCO_3 used in the reaction and the experimental mass of Na_2CO_3 that was produced. **You need to show work in the "calculation" table for how you got your answer. Typing it out is okay for the calculation table section.**
3. Answer the 2 conclusion questions at the end.

Data table

Mass of crucible (g)	12.74 g
Mass of crucible and NaHCO_3 (g)	14.33 g
Mass of crucible and Na_2CO_3 (g)	13.75 g

Calculations Table (show work in each box)

Mass of NaHCO_3 (g)	1.59 g
Experimental (actual) mass of Na_2CO_3 (g)	1.01 g

Conclusion Questions: Show all work, handwritten, using the correct steps/methods shown during class. You will need to upload an image of your work.

1. Use your starting mass of NaHCO_3 and the balanced chemical equation above to calculate the mass of Na_2CO_3 that **should** be produced by the decomposition reaction. This is the theoretical yield.

Baking Soda Stoichiometry Lab Answers

David Callaghan



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