

Ideal Gas Law

Chem Worksheet 14-4

Name _____

The ideal gas law is an equation that relates the volume, temperature, pressure and amount of gas particles to a constant. The ideal gas constant is abbreviated with the variable R and has the value of $0.0821 \text{ atm}\cdot\text{L}/\text{mol}\cdot\text{K}$. The ideal gas law can be used when three of the four gas variables are known. When using this equation it is important that the units for **pressure** are atmospheres (atm), **volume** is in liters (L), and **temperature** is converted to kelvins (K). The **amount** of gas is measured in units called moles (mol).

USEFUL EQUATIONS

$$PV = nRT$$

$$R = 0.0821 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}}$$

$$T_K = T_C + 273$$

$$1 \text{ cm}^3 = 1 \text{ mL}$$

$$1 \text{ L} = 1000 \text{ mL}$$

$$1.00 \text{ atm} = 101300 \text{ Pa}$$

$$1.00 \text{ atm} = 101.3 \text{ kPa}$$

$$1.00 \text{ atm} = 760 \text{ mmHg}$$

$$1.00 \text{ atm} = 760 \text{ torr}$$

$$1.00 \text{ atm} = 14.7 \text{ psi}$$

Unknown	Equation	Known Variables
pressure	$P = \frac{nRT}{V}$	amount, temp., volume
volume	$V = \frac{nRT}{P}$	amount, temp., pressure
temperature	$T = \frac{PV}{nR}$	pressure, volume, amount
amount	$n = \frac{PV}{RT}$	pressure, volume, temp.

example

The pressure exerted by 2.8 moles of argon gas at a temperature of 85°C is 420 torr. What is the volume of this sample?

- list the variables: $P = 420 \text{ torr}$ $V = ?$ $n = 2.8 \text{ mol}$ $R = 0.0821 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}}$ $T = 85^\circ\text{C}$
- convert the variables: $\frac{420 \cancel{\text{torr}}}{1} \times \frac{1 \text{ atm}}{760 \cancel{\text{torr}}} = 0.553 \text{ atm}$ $T = 85^\circ\text{C} + 273 = 358 \text{ K}$
- substitute into the equation: $V = \frac{nRT}{P}$ $V = \frac{(2.8 \cancel{\text{mol}})(0.08206 \frac{\text{atm}\cdot\text{L}}{\text{mol}\cdot\text{K}})(358 \text{ K})}{0.553 \cancel{\text{atm}}} = 82 \text{ L}$

Solve the following problems.

1. A tank contains 115 moles of neon gas. It has a pressure of 57 atm at a temperature of 45°C . Calculate the volume of the tank.
2. A scuba tank has a pressure of 195 atm at a temperature of 10°C . The volume of the tank is 350 L. How many moles of air are in the tank?
3. A helium-filled balloon has a volume of 208 L and it contains 9.95 moles of gas. If the pressure of the balloon is 1.26 atm, determine the temperature in Celsius degrees.
4. A tank of oxygen has a volume of 1650 L. The temperature of the gas inside is 35°C . If there are 9750 moles of oxygen in the tank what is the pressure in PSI?
5. A canister of acetylene has a volume of 42 L. The temperature of the acetylene is 305 K and the pressure is 780 torr. Determine the amount (moles) of gas in the canister.
6. Calculate the volume of a CO_2 cartridge that has a pressure of 850 PSI at a temperature of 21°C . The cartridge contains 0.273 mol of CO_2 .
7. A tank contains 2500 L of argon gas. The pressure is 13790 kPa and the temperature is 25°C . How many moles of argon are in the tank?

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