

Chemical Reactions Primer

Observation and Classification

Introduction

Chemical reactions are the lifeblood of chemistry. Recognizing chemical reactions and “translating” them into chemical equations are essential skills that students must develop in order to be successful in chemistry. Use the following activity to give students extra practice observing and classifying chemical reactions or to assess student understanding of the five basic types of reactions.

Concepts

- Chemical reactions
- Combination vs. decomposition
- Single vs. double replacement
- Combustion reactions

Materials

Ammonia (ammonium hydroxide), 6 M, NH_3 , 5 mL

Bromocresol green solution, 0.04% aqueous, 1 mL

Copper(II) chloride solution, 1 M, CuCl_2 , 15 mL

Ethyl alcohol, $\text{C}_2\text{H}_5\text{OH}$, 1 mL

Hydrochloric acid, 3 M, HCl , 10 mL

Hydrogen peroxide solution, 8%, H_2O_2 , 5 mL

Iron(III) chloride solution, 1 M, FeCl_3 , 5 mL

Iron wire, Fe, 4–5 cm

Phenol red solution, 0.02% aqueous, 1 mL

Sodium hydroxide solution, 1 M, NaOH , 5 mL

Sodium phosphate solution, 0.5 M, Na_3PO_4 , 5 mL

Sulfur powder, S, 0.1 g

Sulfuric acid, 1 M, H_2SO_4 , 2 mL

Water, distilled

Yeast, Baker's, 0.3 g

Zinc, reagent, Zn, 5–6 small pieces

Barrier burner

Butane safety light

Deflagration spoon

Erlenmeyer flask, 250-mL

Evaporating dish, porcelain

Evacuated cylinders, 10-mL, 5

Heat-resistant pad

Pipets, Beral-type, 5

Spatula

Test tubes, 18 × 150 mm, 7

Test tube clamp

Test tube rack

Wood splints, 2

Safety Precautions

Ethyl alcohol is a flammable solvent and a dangerous fire risk. Keep away from flames and other sources of ignition. Solvent bottles should be kept capped at all times and must be removed from the work area when using the laboratory burner. Concentrated ammonia solution is extremely corrosive and toxic by inhalation. Ammonia vapors are extremely irritating to the eyes and respiratory tract. Sulfuric acid, hydrochloric acid and 3 M hydrochloric solutions are corrosive liquids. Keep sodium carbonate and citric acid on hand to clean up acid and base spills, respectively. Hydrogen peroxide solution is a strong oxidizing agent and is corrosive to skin, eyes and the respiratory tract. Sulfur powder is flammable, and dried yeast may also contain a flammable dust. Burning sulfur generates sulfur dioxide, a toxic gas with a stinging, irritating odor. Work with burning sulfur in the hood or in a well-ventilated lab only. Copper(II) chloride solution is slightly toxic by ingestion. Iron(III) chloride and sodium phosphate solutions may be irritating to the eyes and skin. Avoid contact of all chemicals with eyes and skin. Wear chemical splash goggles and chemical-resistant gloves and apron. Please review current Safety Data Sheets for additional safety, handling and disposal information.

Procedure

Record the color and appearance of the reactant(s), the evidence for the chemical reaction and the properties of the product(s). (Optional) Ask students to set up a data table to record observations, and help them interpret the reactions when the activity is over.

Reaction 1

1. Measure 50 mL of distilled water into a 250-mL Erlenmeyer flask, and add about 10 drops of bromocresol green indicator. The water will be green or blue at this point.
2. Fill a deflagration spoon about 1/4 full with powdered sulfur. *Caution:* Carry out steps 3–5 in a fume hood or a well-ventilated lab only.
3. Heat the spoon in a laboratory burner flame until the sulfur melts, turns red and starts to burn.
4. Holding the deflagration spoon upright, lower the spoon into the 250-mL Erlenmeyer flask until the spoon rests about 1–2 cm above the water surface in the flask. Do not allow the spoon to touch the water.
5. Observe the resulting color change in the indicator solution as the gaseous reaction product(s) dissolves in water.

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