

- ④ To find the molecular formula, multiply the empirical formula by the ratio of molar mass divided by empirical mass. In this example it is $66.0/33.0$, which equals 2.
- 4 a Calcium ions are $2+$ and chloride ions are $1-$, so the formula is CaCl_2 .
 b Silver ions are $1+$ and nitrate ions are $1-$, so the formula is AgNO_3 .
 c Copper(II) is Cu^{2+} and phosphate is PO_4^{3-} . For a neutral compound you need three copper ions and two phosphate ions, making charges of $+6$ and -6 . Therefore, the formula is $\text{Cu}_3(\text{PO}_4)_2$.
 d Aluminium is $3+$ and oxide is $2-$, so aluminium oxide is Al_2O_3 .
 ⑤ There are two ways of working out formulae of ionic compounds. One way is to make sure that the total positive charge equals the total negative charge (see part c). The other way, if the charges are not the same, is to swap them round. Copper is $2+$ and phosphate $3-$, so there are three copper ions and two phosphate ions in the formula.
- 5 a As there is one manganese ion to each carbonate ion, the charges have the same numerical value. Carbonate ions are $2-$, so the manganese ions in this compound must be $2+$.
 b The charge on two vanadium ions must equal the charge on three sulfate ions. Each sulfate is -2 , so three sulfate ions have a total charge of -6 . This means that two vanadium ions have a total charge of $+6$, so each has a charge of $+3$.
- 6 a molar mass of $\text{Ca}(\text{OH})_2 = 40.1 + 2 \times (16.0 + 1.0) = 74.1 \text{ g mol}^{-1}$
 b molar mass of $\text{Al}_2(\text{SO}_4)_3 = (2 \times 27.0) + 3 \times [32.1 + (4 \times 16.0)] = 342.3 \text{ g mol}^{-1}$
 c molar mass of $\text{FeSO}_4 \cdot 7\text{H}_2\text{O} = 55.8 + 32.1 + (4 \times 16.0) + 7 \times (2.0 + 16.0) = 277.9 \text{ g mol}^{-1}$
 ⑥ In calculating molar mass, make sure that you use the relative atomic masses (the larger number), not the atomic numbers, from the periodic table. Molar mass has the units of g mol^{-1} . Relative molecular mass has no units. The numbers are the same.
- 7 The balanced equations are:
 a $\text{P}_4 + 5\text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$ or $2\text{P}_2\text{O}_5$
 ⑥ Phosphorus(V) oxide exists as the dimer P_4O_{10} , but $2\text{P}_2\text{O}_5$ on the right of the equation would be acceptable.
 b $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 c $2\text{Mg}(\text{NO}_3)_2 \rightarrow 2\text{MgO} + 4\text{NO}_2 + \text{O}_2$
 ⑥ In part c, balance the oxygen atoms last. Remember that the number of oxygen atoms on the right-hand side of the equation must be an even number, because $\text{Mg}(\text{NO}_3)_2$ contains an even number of oxygen atoms.
 d $2\text{LiOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Li}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 e $2\text{Fe}^{2+} + 2\text{I}^- \rightarrow 2\text{Fe}^{3+} + \text{I}_2$
 ⑥ Make sure that this equation balances for charge. Having one Fe^{2+} ion on the left-hand side and one Fe^{3+} on the right-hand side balances the number of Fe ions, but the charge on the left would not equal the charge on the right.
 f $2\text{C}_6\text{H}_{14} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$ or $\text{C}_6\text{H}_{14} + 12\frac{1}{2}\text{O}_2 \rightarrow 8\text{CO}_2 + 9\text{H}_2\text{O}$
 ⑥ The only source of oxygen on the left-hand side of the equation is O_2 , so the number of oxygen atoms on the right-hand side must be an even number. Therefore, there must be an even number of H_2O molecules on the right. This requires two C_6H_{14} molecules on the left.
- 8 For equations a–c, the simplest way is to realise that the component ions of the compound combine to give the precipitate. All other ions are spectators.

Teaching Transparency Chemistry Answer Key Chapter 19

Muhammad Ishfaq



Teaching Transparency Chemistry Answer Key Chapter 19:

Chapter Resource 2 Chemistry of Life Biology Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004

Resources in Education, 1994 Glencoe Earth Science Ralph M. Feather, 1999 Earth science is the study of Earth and space. It is the study of such things as the transfer of energy in Earth's atmosphere, the evolution of landforms, patterns of change that cause weather, the scale and structure of stars, and the interactions that occur among the water, atmosphere, and land. Earth science in this book is divided into four specific areas of study: geology, meteorology, astronomy, and oceanography. p. 89

Computational Methods in Medicinal Chemistry, Pharmacology, and Toxicology Muhammad Ishfaq, 2025-06-13

Computational Methods in Medicinal Chemistry, Pharmacology, and Toxicology is a comprehensive resource that offers an advanced overview of computational techniques employed in drug discovery, design, and toxicity prediction. The book discusses various topics, including molecular modeling, virtual screening, machine learning, and network pharmacology. It serves as an essential guide for researchers, practitioners, and students in pharmacology, toxicology, medicinal chemistry, bioinformatics, and systems biology fields, showcasing practical applications and future perspectives on new technologies. In addition to covering computational approaches, the book provides real-world examples of drug discovery, candidate optimization, and safety assessment. Other sections explore computer applications in pharmacology and toxicology and discuss the importance of these methods in advancing medicinal research. Offers comprehensive coverage of computational methods that are relevant to pharmacology and toxicology, including molecular modeling, virtual screening, machine learning, and network pharmacology. Includes practical examples and case studies that demonstrate how these methods can be applied in drug discovery, design, and toxicity prediction. Discusses emerging trends and future directions in the field of computational pharmacology and toxicology that can help readers stay up to date with the latest advances and anticipate future developments.

ENC Focus, 2003 **Intelligently Designed** Edward Caudill, 2013-11-15

Tracing the growth of creationism in America as a political movement, this book explains why the particularly American phenomenon of anti-evolution has succeeded as a popular belief. Conceptualizing the history of creationism as a strategic public relations campaign, Edward Caudill examines why this movement has captured the imagination of the American public from the explosive Scopes trial of 1925 to today's heated battles over public school curricula. Caudill shows how creationists have appealed to cultural values such as individual rights and admiration of the rebel spirit, thus spinning creationism as a viable, even preferable, alternative to evolution. In particular, Caudill argues that the current anti-evolution campaign follows a template created by Clarence Darrow and William Jennings Bryan: the Scopes trial's primary combatants. Their celebrity status and dexterity with the press prefigured the Moral Majority's 1980s media blitz. More recent, staunchly creationist politicians such as Sarah Palin and Mike Huckabee, and creationists' savvy use of the Internet and museums to publicize their cause. Drawing from trial transcripts, media sources, films, and archival documents, *Intelligently Designed* highlights the

importance of historical myth in popular culture religion and politics and situates this nearly century old debate in American cultural history

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Introductory Chemistry Nivaldo J. Tro,2006

Essentials of Chemical Education Hans-Dieter Barke,Günther Harsch,Siegbert Schmid,2011-10-28 For everybody teaching chemistry or becoming a chemistry teacher the authors provide a practice oriented overview with numerous examples from current chemical education including experiments models and exercises as well as relevant results from research on learning and teaching With their proven concept the authors cover classical topics of chemical education as well as modern topics such as every day life chemistry student s misconceptions the use of media or the challenges of motivation This is the completely revised and updated English edition of a highly successful German title

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